

Gold transport in hydrothermal chloride-bearing fluids: Insights from in situ x-ray absorption spectroscopy and ab initio molecular dynamics

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Originally published:

December 2018

ACS Earth and Space Chemistry 3(2019)2, 240-261

DOI: <https://doi.org/10.1021/acsearthspacechem.8b00103>

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27 **ABSTRACT**

28 Chloride-rich fluids are widespread in the Earth's interior from low-temperature subsurface
29 conditions to deep lithosphere. The concentration of chloride salts vary from diluted aqueous
30 solutions peculiar to epithermal deposits to concentrated chloride brines with salinity up to 70 wt%
31 NaCl eq., which coexist with low density vapor during the formation of porphyry systems, and
32 anhydrous (dry) chloride melts at the upper part of degassing magma chamber beneath volcanoes.
33 Here we report an investigation of the speciation of Au in hydrothermal chloride fluids and
34 anhydrous melt by means of X-ray absorption spectroscopy (XAS). The experiments consisted in
35 registration of Au L_3 -edge X-ray absorption near edge structure/extended X-ray absorption fine
36 structure (XANES/EXAFS) spectra of Au-bearing fluids in the temperature range from 350 to 575
37 °C at pressures 0.5-4.5 kbar. The experimental aqueous systems consisted of HCl-, NaCl-, KCl-, or
38 CsCl- dominated fluid (0.02 to $7 \text{ mol} \cdot (\text{kg H}_2\text{O})^{-1}$) inside a silica glass capillary together with small
39 piece of Au, and SO_2/SO_3 buffer or $\text{K}_2\text{S}_2\text{O}_8$ as a source of oxygen. Spectra of Au dissolved in dry
40 CsCl/NaCl/KCl + $\text{K}_2\text{S}_2\text{O}_8$ melt were recorded at 650 °C. The measured concentrations of Au
41 dissolved in hydrothermal fluids were in close agreement with the values calculated using
42 thermodynamic data available for AuCl_2^- . High concentration of Au dissolved in dry chloride melt
43 implies that it can be an effective Au scavenger in natural systems with low H_2O content. It was
44 determined that Au is coordinated by two Cl atoms ($R_{\text{Au-Cl}} = 2.25 - 2.28 \text{ \AA}$) and 0.2 – 0.6 alkali
45 metal atoms (Me) in the distant coordination sphere of Au at $R_{\text{Au-Me}} = 3.3 - 4.1 \text{ \AA}$. The presence of
46 the alkali metal cations in the vicinity of Au-Cl complex was confirmed by the XANES spectra
47 registered in High Energy Resolution Fluorescence Detection (HERFD) mode: the white line (WL)
48 intensity decreased substantially in NaCl-dominated fluid and dry melt in comparison with HCl-
49 dominated fluid. The observed decrease of the WL intensity stems from the partial compensation of
50 the positive charge located on Au due to the presence of the distant-coordination-sphere alkali metal
51 cations. The interatomic distances between Au and atoms localized in the first and distant
52 coordination spheres exhibit contrasting behavior with respect to fluid pressure. Increasing pressure

53 results in weak contraction of the first coordination sphere of Au, whereas conjugated increase of
54 the dielectric constant causes expansion of the distant coordination sphere. These results can be
55 explained by the difference in the chemical bond nature. The Cl⁻ ions in the nearest coordination
56 sphere of Au are bonded to the central atom via the strong (predominantly covalent) chemical
57 bonding, whereas the distant-coordination-sphere cations are weakly bonded to the negatively
58 charged AuCl₂⁻ complex mostly via the electrostatic interaction. This means that in thermodynamic
59 calculations the dissolution of Au in natural ore-forming fluids, at least of moderate to high
60 densities (> 0.5 g·cm⁻³), can be explicitly described by reactions which involve the formation of
61 AuCl₂⁻, with mandatory consideration of the effect of weak outer sphere coulomb interactions via
62 the calculation of activity coefficients.

63 1. INTRODUCTION

64 Chlorides are the most abundant salt components of hydrothermal fluids in the Earth's crust.
65 The concentration of NaCl, the main electrolyte salt, varies in wide limits, with the average value
66 close to $\sim 0.5\text{-}1\ m$ (see, for example, review of hydrothermal fluid compositions in Seward et al.,
67 2014). In particular, in hydrothermal fluids which formed the volcanogenic massive sulfide (VMS)
68 deposits the chloride concentration ranged from few to 20 wt% NaCl eq. (e.g., Large, 1977;
69 Simonov et al., 2006; Vikent'ev et al., 2009, 2012; Karpukhina et al., 2013). Similar concentrations
70 of chloride salts, with the average value of 4-8 wt% NaCl eq. and maximum concentration up to 26
71 wt% NaCl eq., were established in hydrothermal vent fluids of black smokers – modern analogues
72 of VMS deposits (Douville et al., 2002; Bortnikov et al., 2004, 2014; Schmidt et al., 2011).
73 Chlorinity of the high-temperature ($t\sim 600\text{-}800\ ^\circ\text{C}$) metal-rich fluid rising from a magma chamber in
74 porphyry systems is about 10 wt% NaCl eq. Lowering of the temperature and pressure induces
75 phase separation that results in formation of a low density fluid with salinity of a few wt% NaCl eq.
76 and a brine with salinity of 30-70 wt% NaCl eq. (Heinrich et al., 2004; Sillitoe, 2010; Gregory,
77 2017). Note that many VMS and porphyry systems belong to the World-class gold deposits with Au
78 resources of more than 100 t. Gold alloys with platinum group elements (PGE: Pt, Pd, Ir, and Rh)
79 are frequent in giant PGE deposits of magmatic origin - deposits of Norilsk, Bushveld, and
80 Stillwater. In all these deposits hydrothermal fluids participated in the formation of PGE-Au
81 mineralization. In the Stillwater Complex the concentration of chloride salts ranged from NaCl-
82 dominated halide melts ($>82\ \text{wt}\% \text{ NaCl eq.}, t = 660\text{-}800^\circ\text{C}$) to more complex Na-Ca-K-Fe-Mn-Ba
83 brines of lower salinity ($28\text{-}79\ \text{wt}\% \text{ NaCl eq.}, t = 480\text{-}640\ ^\circ\text{C}$, Hanley et al., 2008). Chloride-rich
84 fluids discovered as fluid inclusions in quartz associated with sulfide minerals at Merensky reef
85 (Bushveld complex) were entrapped in the temperature range $750\text{-}400\ ^\circ\text{C}$. Their salinity varies from
86 high ($60\text{-}70\ \text{wt}\% \text{ NaCl eq.}$) to moderate-to-low ($20\ \text{to}\ \sim 7\ \text{wt}\% \text{ NaCl eq.}$) depending on the
87 entrapment temperature (Ballhaus and Strumpfl, 1986; Zhitova et al., 2016). Abundance of H_2O -,
88 Cl -, and F - bearing minerals crystallized in Norilsk ores in the vicinity of Au chalcogenides,

89 chlorides, and intermetallic compounds suggested that the Au mineralization was related to the
90 post-magmatic hydrothermal stage which took place at temperature of ~350 °C and below
91 (Sluzhenikin and Distler, 2015).

92 Native gold crystals and particles are often formed during volcanic eruptions and fumarolic
93 activity. Crystals of native gold were found in products of volcanic eruptions which were formed
94 during 1975-1976 and 2012-2013 eruptions of Tolbachik volcano, Kamchatka (see Chaplygin et al.,
95 2015, and references cited). Micron size native gold particles were discovered on the surface of
96 sulfide minerals (wurtzite-greenockite, (Zn,Cd)S) formed in active fumaroles of Kudriavy volcano,
97 Iturup Island (Chaplygin et al., 2007). High-temperature fumarole gases at these volcanoes,
98 contained both Au (few ppb to few tens ppb) and chloride (as high as 1.2 mol% HCl at Tolbachik
99 volcano, Chaplygin et al., 2015). Association of Au-bearing phases with oxychlorides and chlorides
100 observed in tube precipitates collected at Kudriavy volcano also suggests elevated Au and Cl
101 concentrations in the gas phase (Yudovskaya et al., 2006). Almost dry (anhydrous) chloride salt
102 melt as a reservoir which accumulates metals at the upper part of degassing magma chamber
103 beneath a volcano was suggested by Shmulovich and Churakov (1998) to explain the formation of
104 the metal-rich gas sublimates. High capacity of chloride melts to accumulate Au is demonstrated by
105 their effective usage as a salt flux for the synthesis of crystals (cf. Chareev et al., 2016) and is
106 confirmed in the experiments on the dissolution of Au in (Na,K)Cl melts (Shmulovich et al., 2016).
107 The speciation of Au in melted chloride salts, as well as the dissolution mechanism is, however,
108 unknown.

109 The chemistry of Au in chloride systems has been experimentally studied via measurements
110 of Au solubility at 300-600 °C and pressure up to 1800 bar (e.g., Zotov and Baranova, 1989; Zotov
111 et al., 1989, 1991, 1995, 2017; Archibald et al., 2001; Stefánsson and Seward, 2003; Guo et al.,
112 2018, and references therein). It was determined that in chloride-bearing acidic fluids Au can be
113 effectively transported in the form of Au-Cl complexes with predominant role of AuCl_2^- , even
114 though some inconsistencies still remain regarding the complex stability. The Au speciation model

115 was confirmed by X-ray absorption spectroscopy (XAS) (Pokrovski et al., 2009). Mei et al. (2014)
116 studied the complexing of Au with chloride in NaCl-rich aqueous fluids by means of *ab initio*
117 molecular dynamics (AIMD) simulations performed at conditions from ambient to 1000 °C, 1500
118 bar. It was determined that in concentrated NaCl solutions AuCl_2^- is associated with Na^+ ion which
119 is weakly bonded in the outer coordination sphere. The authors proceeded with the calculation of
120 thermodynamic properties of $\text{NaAuCl}_2^\circ_{(\text{aq})}$ complex and found that its stability rapidly increases
121 with decrease of the hydrothermal fluid density. These data imply that neutral ionic associates
122 predominate at relatively low densities ($d < 0.7 \text{ g/cm}^3$), whereas charged complexes dominate the
123 Au speciation in higher density fluids. At the same time, the dissolved Au concentration measured
124 in the cited above Au solubility studies in the wide range of NaCl/KCl concentrations and fluid
125 densities, including $d < 0.7 \text{ g/cm}^3$, can be accurately described without the formation of
126 $\text{Na(K)AuCl}_2^\circ_{(\text{aq})}$. In the works of Frank et al. (2002) and Zajacz et al. (2010) the solubility of Au
127 was measured in high-temperature fluids of contrast densities (dense brines with total chloride
128 content of 70 wt% at 800°C, and low density fluids at 1000 °C/1500 bar, respectively). Positive
129 correlations between HCl and NaCl/KCl concentrations with the solubility of Au was observed and
130 interpreted in terms of the formation of HAuCl_2° (Frank et al., 2002) and NaAuCl_2° (Zajacz et al.,
131 2010). However, complex system compositions and the absence of the reliable thermodynamic data
132 on the system's components did not allow the authors to determine the dominant Au complex
133 composition and calculate its stability constant unambiguously.

134 The aim of this work is to quantify the impact of temperature, pressure, and the system
135 compositions on the chemistry of Au in chloride-rich hydrothermal fluids by means of synchrotron-
136 based XAS. Gold L_3 -edge X-ray absorption near edge structure/extended X-ray absorption fine
137 structure (XANES/EXAFS) spectra were recorded in situ for HCl-, NaCl-, KCl-, and CsCl-
138 dominated fluids in the temperature range from 350 to 575 °C at pressures of 0.5-4.5 kbar. The X-
139 ray absorption spectra of dry CsCl/NaCl/KCl melt were acquired at 650 °C. Most of the
140 experiments were performed using the Total Fluorescence Yield (TFY) detection method while the

141 XANES spectra of two solutions (NaCl- and HCl- dominated), and the dry chloride melt were
142 registered in the High Energy Resolution Fluorescence Detection (HERFD) mode (Glatzel and
143 Bergman, 2005). The presence of cation was detected in the second coordination sphere of Au for
144 alkali metal – dominated fluids and melts. Despite the fact that geometry of the distant coordination
145 sphere is highly disordered, we found that the outer-sphere cation affects the charge state of Au.
146 The contrast effect of pressure was determined on interatomic distances between Au and atoms
147 localized in its nearest and distant coordination spheres.

148

149 **2. METHODS**

150 **2.1. X-ray absorption spectroscopy**

151 X-ray absorption experiments were performed at the European Synchrotron Radiation
152 Facility (ESRF) in Grenoble, France. The Au L_3 -edge TFY XAS (XANES/EXAFS) spectra were
153 recorded at the Rossendorf Beamline BM20 (ROBL). The storage-ring operating conditions were
154 6.0 GeV and 80-100 mA. The photon energy was scanned from 11620 to 12670 eV using the
155 Si(111) monochromator coupled to Rh-coated mirrors for the collimation and reduction of higher
156 harmonics. Energy calibration was performed using the L_3 -edge excitation energy of Au metal foil
157 (11919 eV). The TFY XAS spectra were collected using 13-element high-throughput Ge-detector.
158 The total energy resolution (incident energy and core – hole lifetime broadening) has been
159 evaluated as 8.8 eV. The detected intensity was normalized to the incident photon flux.

160 The Au L_3 -edge HERFD-XANES spectra were collected at the high-brilliance X-ray
161 absorption/X-ray emission spectroscopy undulator beamline ID26 (Gauthier et al. 1998). The ring
162 current was varied between 150 and 200 mA. The incident energy was selected using the $\langle 111 \rangle$
163 reflection from a double Si crystal monochromator. Rejection of higher harmonics was achieved by
164 three Pd mirrors positioned at an angle of 2.5 mrad relative to the incident beam. The incident X-ray
165 beam had a flux of approximately $2 \cdot 10^{13}$ photons s^{-1} on the sample position. The spectra were
166 measured in the HERFD mode using an X-ray emission spectrometer (Glatzel and Bergman 2005;

167 Kvashnina and Scheinost 2016). The sample, analyzer crystal and photon detector (silicon drift
168 diode) were arranged in a vertical Rowland geometry. The Au L_3 -edge HERFD-XAS spectra were
169 obtained by recording the intensity of the Au $L_{\alpha 1}$ emission line (9713 eV) as a function of the
170 incident energy. The emission energy was selected using the $\langle 555 \rangle$ reflection of four spherically
171 bent Ge crystal analyzers (1 m curvature radius) aligned at 78° Bragg angle. A combined (incident
172 convoluted with emitted) energy resolution of 1.5 eV was determined as the full width at half
173 maximum of the elastic peak. The intensity was normalized to the incident flux.

174 The experimental set-up used for the high-temperature in situ XAS experiments is described
175 in detail in Trigub et al. (2017a,b) and will be only briefly outlined here. The heating experiments
176 were performed using a capillary technique initially developed for Raman spectroscopic studies
177 (Chou et al., 2008, 2012; Ding, 2010). A small piece of Au wire (1.2-1.5 mm long) was loaded into
178 a Polymicro TechnologiesTM silica glass capillary: 600 μm OD, 250 μm ID, 12 mm length (BM20
179 beamline), or 600 μm OD, 400 μm ID, 15 mm length (ID26 beamline). The capillaries were filled
180 with experimental solutions and hermetically sealed (the capillaries were water-cooled during the
181 sealing). The redox state of the system was controlled either by sulfite/sulfate equilibrium (a
182 solution with known concentration of $\text{SO}_2/\text{H}_2\text{SO}_4$ or $\text{Na}_2\text{SO}_3/\text{H}_2\text{SO}_4$ was loaded into the capillary),
183 or by oxygen produced due to the thermal decomposition of $\text{K}_2\text{S}_2\text{O}_8$ ($\text{K}_2\text{S}_2\text{O}_8 \rightarrow \text{K}_2\text{SO}_4 + \text{SO}_2 + \text{O}_2$). In
184 the latter case a weighted amount (~ 0.02 - 0.2 g) of $\text{KCl} + \text{K}_2\text{S}_2\text{O}_8$ mixture was loaded into the
185 capillary prior to an experimental solution. The HCl concentration was determined by volumetric
186 titration against Trizma[®] base using methyl red as indicator, the concentration of SO_2 - by
187 iodometric titration, and the concentration of H_2SO_4 by densimetry. In case of dry chloride melt the
188 capillaries were loaded with the eutectic mixture of $\text{CsCl}/\text{NaCl}/\text{KCl}$ (30 at% $\text{NaCl}/24.5$ at% KCl
189 with melting temperature of 478°C (Fullam, 1971) together with $\text{K}_2\text{S}_2\text{O}_8$. The volumetric ratio of
190 the chloride mixture to $\text{K}_2\text{S}_2\text{O}_8$ was $\sim 10/1$. The capillaries with eutectic chloride mixture were
191 evacuated down to a 10^{-4} bar pressure, sealed, and preconditioned at the experimental temperature
192 during 24 hours. At the beamline a capillary was placed into the microtomography furnace (Bellet

193 et al., 2003) between two plates of silver (BM20 beamline) or stainless steel (ID26 beamline) which
 194 eliminated temperature gradient. Temperature readings were calibrated before the experiments at an
 195 accuracy of ± 5 °C with a *K*-type thermocouple inserted directly in the place of the capillary.
 196 Pressure inside the capillary was estimated from the degree of filling of the capillary using *PVT*
 197 properties of H₂O-NaCl system (Driesner, 2007; Driesner and Heinrich, 2007) or, in case of HCl-
 198 dominated fluids, H₂O-HCl (Bach et al., 1977). We estimate the uncertainty of the calculated
 199 pressures as $\pm 10\%$ for HCl- and NaCl-dominated fluids, and $\pm 25\%$ for KCl- and CsCl- dominated
 200 ones. The spectra were recorded after thermal equilibration of the experimental system and the
 201 signal stabilization. Usually ca. 30-45 min at the experimental temperature was necessary to attain
 202 the stationary state. The parameters of XAS experiments and the experimental system compositions
 203 are listed in Table 1.

204

205 **2.2. EXAFS spectra fitting and wavelet transforms**

206 The EXAFS ($\chi_{exp}(k)$) data were analyzed using ARTEMIS program (a part of IFEFFIT
 207 software package, Ravel and Newville, 2005). Following standard procedures for pre-edge
 208 subtraction and spline background removal, the structural parameters - interatomic distances (R_i),
 209 coordination numbers (N_i), and Debye–Waller factors (σ_i^2) - were determined via the non-linear fit
 210 of the theoretical spectra to the experimental ones with the equation

$$211 \quad \chi(k) = S_0^2 \sum_{i=1}^n \frac{N_i F_i(k)}{R_i^2 k} e^{\frac{-2R_i}{\lambda(k)}} e^{-2\sigma_i^2 k^2} \sin(2kR_i + \varphi_i(k)) \quad . \quad (1)$$

212 The theoretical spectra were simulated using photoelectron mean free path length $\lambda(k)$, amplitude
 213 $F_i(k)$, and phase shift $\varphi_i(k)$ parameters calculated *ab initio* using the program FEFF6 (Zabinsky et
 214 al., 1995).

215 Wavelet transform (WT) analysis of the EXAFS spectra was performed as described in
 216 Penfold et al. (2013), Timoshenko and Kuzmin (2009), Funke et al. (2007). The WT consists in
 217 replacement of the infinitely expanded periodic function in Fourier transform by a local function -

218 the wavelet. The WT allows analyzing the experimental EXAFS spectra in two-dimensional k/R
219 space and enables separation of the contributions from atoms with different Z . The WT of a given
220 signal $\chi(k)$ is defined as

$$221 \quad W_f^\psi(a, k') = \frac{1}{\sqrt{a}} \int_{-\infty}^{+\infty} \chi(k) \psi^* \left(\frac{k - k'}{a} \right) dk, \quad (2)$$

222 where the scalar product of the EXAFS signal and the complex conjugate of the wavelet (ψ^*) is
223 calculated as a function of a and k' . The a parameter is related with R as $a = \frac{\eta}{2R}$, and k' conforms
224 to localization of the wavelet function in k -space. In this work we used the WT based on the Morlet
225 wavelet functions

$$226 \quad \psi(k) = \frac{1}{\sqrt{2\pi\sigma}} e^{i\eta k} e^{-k^2/2\sigma^2}, \quad (3)$$

227 where parameters σ and η are width and frequency of the wavelet function, respectively. These
228 parameters should be adjusted to get appropriate resolution in k - and R -space. For providing better
229 quality of WT images we used modified WT functions (Timoshenko and Kuzmin, 2009).

230

231 **2.3. *Ab initio* molecular dynamics (AIMD) simulations**

232 The AIMD simulations are aimed to facilitate the interpretation and modeling of XAS
233 results. Details of the AIMD simulations are described in Trigub et al. (2017b). Born–Oppenheimer
234 AIMD simulations were performed with the aid of CP2K program package (VandeVondele et al.,
235 2005). A hybrid Gaussian plane-wave method (GPW) (Lippert et al., 1997) was used. The
236 Goedecker–Teter–Hutter (GTH) norm-conserving pseudo-potentials (Goedecker et al., 1996;
237 Hartwigsen et al., 1998; Krack, 2005) were applied to replace core electrons. The electronic density
238 cut off was set at 500 Ry. For all the atoms the Double-zeta valence plus polarization (DZVP) basis
239 set optimized for molecular geometries was used. The spin polarized general gradient
240 approximation by Perdew–Burke–Erzerhof (PBE) (Perdew et al., 1996) augmented with a
241 correction of the van der Waals (vdW) interactions was applied as exchange correlation functional,

242 the nonlocal rVV10 vdW functional was used (Vydrov et al., 2009; Sabatini et al., 2013). The b
243 parameter which determines the nonlocal behavior of the functional was set to 9.3 according to
244 Miceli et al. (2015). The simulations were conducted in the canonical NVT ensemble, with the
245 simulation time step of 0.5 fs and periodic boundary conditions. The AIMD trajectory was
246 accumulated for 20 ps and 10 ps for solution and melt, respectively. Each AIMD calculation was
247 preceded by equilibration phase of 5 ps. The temperature was controlled via a Nosé–Hoover chain
248 thermostat (Nosé, 1984a, 1984b). Starting configurations for AIMD calculations were generated by
249 classical MD method with the aid of GROMACS software (Berendsen et al., 1995). The system
250 contained AuCl_2^- or $\text{Cat}[\text{AuCl}_2]$ complex ($\text{Cat} = \text{H}, \text{Na}, \text{K}, \text{Cs}$) which geometry was optimized with
251 Gaussian 09 suite of programs (Frisch et al., 2016) (the Møller-Plesset Perturbation Theory (MP2)
252 was applied; solvation effects were accounted for by means of the Conductor-like Polarized
253 Continuum Model (CPCM), see Pokrovski et al. (2009) for calculation details). Details of the
254 chemical system compositions, box sizes, and temperatures used in the simulations are given in
255 Table 2.

256 An additional calculation was performed to check the effect of basis set size on results of the
257 AIMD calculation. Calculation H (Table 2) was identical to calculation E in terms of the atomic
258 system parameters (system composition, box size), but was performed with triple-zeta valence basis
259 set (TZV2P) applied for all atoms. This calculation was performed only for one system because the
260 calculation time increased substantially (from 25 s to 40 s for one time step).

261 Results of AIMD calculations are presented in the form of plots of atomic densities and
262 coordination numbers. The radial density $\rho(R)$ of surrounding atoms is related to coordination
263 numbers via expression:

$$264 \quad N(R) = \int 4\pi r^2 \rho(R) dr . \quad (4)$$

265

266 **2.4. EXAFS and XANES spectra simulations using AIMD data**

267 Every 50 fs of AIMD simulation snapshots of atomic configurations were extracted and used
268 to calculate Au L_3 -edge EXAFS/XANES spectra (hereafter the AIMD-EXAFS/AIMD-XANES
269 notation will be used for the calculated spectra). The configuration-averaged Au L_3 -edge EXAFS
270 spectra were calculated using FEFF8.5L program (Ankudinov et al., 1998). The Au-Cl distances
271 were corrected by a factor of 0.97 to account for differences between experimental (EXAFS) and
272 theoretical (AIMD) radii of the first coordination sphere. The multiple scattering (MS) contributions
273 to EXAFS were accounted up to the fifth order within a cluster of 6 Å radius around the absorbing
274 atom. The calculation of the cluster potential was performed for every geometry configuration. The
275 self-consistent complex exchange-correlation Hedin–Lundqvist potential and default values of
276 muffin-tin radii as provided within the FEFF8.5L code were applied. The amplitude of the
277 theoretical EXAFS spectrum was multiplied by the $S_0^2 = 0.84$ factor as determined from the Au₂S
278 EXAFS spectra fitting.

279 Theoretical calculations of the Au L_3 -edge HERFD-XANES spectra were performed using
280 the finite difference method (FDM) by means of FDMNES program (Joly, 2001; Guda et al., 2015).
281 Relativistic self-consistent field (SCF) FDMNES calculations were carried out with the exchange-
282 correlation part of potential in a local density approximation (Hedin and Lundqvist, 1971). The final
283 electronic states were calculated in a full core hole screening. Atomic clusters inside the spheres
284 with radii of 6 Å were chosen for both SCF and FDM XANES calculations. Many-body effects and
285 core-hole lifetime broadening were accounted for by means of arctangent convolution (Bunău and
286 Joly, 2009).

287

288 **2.5. Bader charges calculation**

289 The atomic clusters obtained by means of AIMD simulations were used to study the charge
290 state of Au. First, the electron density of each cluster was calculated by means of Quantum Espresso
291 software package (Giannozzi et al., 2009). At this stage we applied a projector-augmented wave
292 (PAW) method with norm-conserving pseudopotentials and the Perdew-Burke-Ernzerhof (PBE)

293 exchange-correlation functional. The SCF method was applied with a 100 Ry kinetic energy cutoff
294 for the plane waves and an SCF tolerance better than 10^{-9} for the electronic structure calculations.

295 Atomic charges were determined according to quantum theory of atoms in molecules
296 (QTAIM, Bader, 1990, 1991). The local atomic charges of Au were calculated by integrating the
297 charge density within a Bader volume around the atom using Critic2 software (Otero-de-la-Roza et
298 al., 2009, 2014).

299

300 **3. RESULTS AND DISCUSSION**

301 **3.1. Quantitative analysis of XANES spectra and dissolved Au concentrations**

302 Gold L_3 -edge HERFD-XANES spectra are shown in Fig. 1, and the TFY-XANES spectra
303 are presented in Fig. 2a,b. Positions of the main spectral features are given in Table 3. In all
304 considered systems the oxidation state of Au is +1 despite high oxygen fugacity, especially in the
305 experiment with chloride melt. The HERFD-XANES spectra of chloride melt, NaCl- and HCl-rich
306 solutions differ in the height of the white line (WL, the first, most intense feature). The most intense
307 WL was observed in experiment cap28 with HCl-dominated fluid (3m HCl/0.02m NaCl, Fig. 1).
308 The WL's intensity of NaCl-rich fluid in cap30 experiment (3m NaCl) and dry chloride melt
309 (cap5407) are considerably lower. The Au L_3 -edge absorption is related to the $2p$ - $5d$ dipole-allowed
310 transitions: $2p_{3/2} \rightarrow 5d_{5/2}/5d_{3/2}$. Therefore, the WL intensity reflects the number of empty states
311 (holes) in $5d_{5/2}$ and $5d_{3/2}$ orbitals above the Fermi level. Low WL intensity of NaCl-rich fluid and
312 chloride melt indicates the reduced unoccupied density of states (DOS) of Au $5d$ orbitals, or, in
313 other words, larger number of electrons localized at Au $5d$ orbitals for salt-rich systems (cap30 and
314 cap5407) than in the HCl-dominated system (cap28). This fact can be quantitatively explained by
315 interaction of positively charged ions with AuCl_2^- complex in concentrated salt fluid and melt. The
316 presence of positively charged ions in the vicinity of AuCl_2^- complex results in electron transfer
317 from ligands to metal, which gives rise to decrease of the number of holes in Au $5d$ orbitals (or
318 decrease of the positive charge localized on Au) and, therefore, decrease of the WL intensity. The

319 decrease of the positive charge is confirmed by the Bader partial atomic charges calculation which
320 is discussed below (section 3.5).

321 The WL intensities of the TFY XANES spectra of hydrothermal fluids (cap122, cap118,
322 cap117, cap110), all of which were recorded for salt-rich systems, are identical (Fig. 2a). At the
323 same time, the WL of the TFY XANES spectrum of salt melt (cap2156) is more intense than WL of
324 K- and Cs-bearing aqueous fluids. This fact can be explained by more effective compensation of Au
325 positive charge in aqueous fluids which contain high concentrations of heavy alkaline metal cations
326 in comparison with chloride salt melt. Besides, according to the AIMD calculations, the positive
327 charge on Au in chloride melt increases due to the presence of Cl^- ions in the distant coordination
328 sphere at $2.6 - 3 \text{ \AA}$, whereas no Cl^- ions were detected in the second coordination sphere of Au in
329 aqueous fluids (see section 3.3). We note, however, that the effect of fluid composition on the
330 electronic density distribution within the Au complex and on the Au valence state remains open.

331 Figure 2b compares XANES spectra of chloride melt (cap2156) recorded at high and
332 ambient temperature. The spectra are different: the crystallized melt spectrum exhibits several
333 additional features which are absent in the high-temperature spectra. This difference is connected
334 with appearance of long-range atomic order in crystallized melt. The edge jump (e.j.) and WL
335 positions of these spectra are also shifted relative to each other (Table 3).

336 The concentrations of dissolved Au calculated from the amplitude of the absorption edge
337 height of the transmission spectra are listed at the bottom of Table 1. They are in close agreement
338 with Au solubilities calculated using thermodynamic properties of AuCl_2^- reported in Akinfiev and
339 Zotov (2010). Good agreement between results of our experiments and calculated solubilities
340 argues for high quality of the thermodynamic model of Akinfiev and Zotov (2010), as well as the
341 experimental data which were used by the authors to calculate thermodynamic properties of AuCl_2^- .
342 Note that relatively high uncertainty of Au concentrations determined in the present study, which
343 we estimate as ± 0.3 log units, stem from the spherical shape of the capillary, possible shift of the
344 beam from the center of the capillary, uncertainty in the solute concentrations due to weighting of

345 small masses of solid reagents added directly to the capillaries, and due to possible loss of volatile
346 components during the sealing of the capillary. These uncertainties can account for the
347 underestimated experimental Au solubilities compared to those calculated using database of
348 Akinfiev and Zotov (2010).

349 We estimate the concentration of Au dissolved in chloride melt as ~ 1 wt%. This
350 concentration (in mol fraction scale) is ca. 10 times lower than the concentration of oxygen
351 produced by decomposition of potassium persulfate. However, it is more than likely that the
352 equilibrium Au concentration has not been attained due to slow kinetics of Au dissolution in
353 chloride melt: we were not able to detect dissolved Au during one hour XAS experiment performed
354 at 650 °C with capillaries which were not subjected to preconditioning. Therefore, we can conclude
355 that the solubility of Au in chloride melts can be very high, but the quantitative thermodynamic
356 description of the Au solubility needs further experimental examination.

357

358 **3.2. EXAFS spectra fitting**

359 Roughly speaking, temperature, pressure, and the experimental system composition have
360 little effect on the shape of Au L_3 -edge EXAFS spectra and their Fourier transforms (at least at $k <$
361 10 \AA^{-1} and $R < 3 \text{ \AA}$, Fig. 3a and b, respectively). Results of EXAFS spectra fitting performed by
362 means of ARTEMIS computer program are shown in Fig. 4 and collected in Table 4. In general,
363 these data are consistent with the results of X-ray absorption spectroscopy, quantum chemical, and
364 AIMD calculations previously reported in the literature (Pokrovski et al., 2009; Mei et al., 2013). In
365 chloride-bearing fluids two Cl^- ions are the nearest neighbors of Au with $R_{\text{Au-Cl}} = 2.25 - 2.28 \text{ \AA}$.
366 Results of the present study, however, reveal some important features of Au chemistry which were
367 not manifested previously. It follows from data of Table 4 that Au-Cl distance in the first
368 coordination sphere weakly decreases with increasing pressure. These small systematic changes
369 were monitored in all the studied systems. The most pronounced change in the $R_{\text{Au-Cl}}$ of 0.02 \AA was
370 observed between experiments cap118 and cap117 (the pressure increased from 2100 to 4500 bar

371 for fluids of identical composition). The maximum value of $R_{\text{Au-Cl}} = 2.28 \text{ \AA}$ was monitored in dry
372 chloride melt when pressure did not exceeded one bar. Therefore, one can conclude that increase of
373 pressure induces contraction of the first coordination sphere of Au.

374 In all the experiments the presence of a heavy Me atom (Me = K or Cs) was observed in the
375 second coordination sphere of Au (the contribution of Au-Me-Au path to the Fourier transform of
376 EXAFS spectra is shown in Fig. A1). The coordination numbers and Debye-Waller parameters of
377 cations were estimated with high uncertainty, but inclusion of Au-Me-Au path in the fits of the
378 EXAFS spectra improved considerably the fit results in the 2.5 - 4 \AA range (not corrected for phase
379 shift) in comparison with the model without the distant-coordination-sphere cation (Table A1, Fig.
380 A2). When it was possible to calculate the coordination number of Me atoms (N_{Me}) in the second
381 coordination sphere of Au the N_{Me} fell within 0.2 - 0.6. In contrast to the observed contraction of the
382 first coordination sphere of Au with increase of pressure, the opposite effect was monitored for the
383 Me atoms in the second coordination sphere. The increase of the aqueous fluid pressure resulted in
384 the increase of the Au-Me distance. The effect of pressure is of sharp focus in experiments cap118
385 ($R_{\text{Au-Me}} = 3.31 \pm 0.05 \text{ \AA}$, $P = 2100 \text{ bar}$, $d = 0.68 \text{ g}\cdot\text{cm}^{-3}$) and cap117 ($R_{\text{Au-Me}} = 4.14 \pm 0.03 \text{ \AA}$, $P = 4500$
386 bar , $d = 1.04 \text{ g}\cdot\text{cm}^{-3}$). We note here that, despite high uncertainty, decrease of the $R_{\text{Au-Me}}$ is of
387 systematic character. It was observed, at a quantitative level, in all the experimental systems even
388 though the exact change of the $R_{\text{Au-Me}}$ as a function of TP -parameters can not be determined via the
389 experimental spectra fitting.

390 The observed changes of the $R_{\text{Au-Me}}$ can be explained by the effect of dielectric constant on
391 the weakly bound distant-coordination-sphere ion. The decrease of density results in the decrease of
392 the fluid dielectric constant which promotes electrostatic interaction of the negatively charged
393 AuCl_2^- complex and positively charged Me^+ ion. As a result, $R_{\text{Au-Me}}$ decreases with decrease of
394 pressure (density, dielectric constant) due to the enhancement of coulomb interaction.

395

396 **3.3. Results of AIMD simulations**

397 Simulation results are given in Table 2 (calculation parameters), Table 5 (geometry of Au
398 species), and shown in Fig. 5. All the simulations for aqueous fluids were performed for 450 °C in
399 order to compare the effect of the system composition on the Au-Cl complexing and the X-ray
400 absorption spectra, and for 650 °C in the case of chloride melt. The first coordination sphere of Au
401 comprises two chloride atoms in all the simulations but simulation F (chloride melt). In the latter
402 case the presence of Cl⁻ ions at 2.6 - 3 Å induces an asymmetric shape of Cl atomic density function
403 (Fig. 5). The distant coordination sphere of Au is composed of H₂O, Cl, and alkaline metal cations.
404 The atomic density curves for Na, K, and Cs is of diffuse character with maxima located between 3
405 and 6 Å. This shape of the density curves confirms our conclusion based on the analysis of EXAFS
406 spectra about weak electrostatic interaction of AuCl₂⁻ with the distant-coordination-shell cations. In
407 this case the cation does not have a definite position, but the $R_{\text{Au-Me}}$ fluctuates within ~3 - 6 Å range.
408 The $R_{\text{Au-Me}}$ depends on the cation type and increases with increasing cation atomic weight ($R_{\text{Au-Na}} =$
409 2.8 - 4.6 Å, $R_{\text{Au-Cs}} = 3.3 - 6$ Å). The coordination number N_{Me} is not a constant but increases with
410 increasing cationic weight and salt concentration (Table 5), which is unlike for a stable aqueous
411 complex which has a definite composition. Computation G, which was performed for the same
412 system composition as the computation E, but with extended atomic basis set, yielded similar
413 results. In both calculations Cs atoms fluctuate at a distance of 3 - 6 Å with the maximum of atomic
414 density at ~ 4 Å. Therefore, the DZVP basis set is accurate enough for AIMD investigation of
415 hydrothermal Au complexing.

416

417 **3.4. AIMD-EXAFS spectra and comparison with experimental data**

418 EXAFS spectra simulated using atomic configurations produced by AIMD calculations are
419 shown in Fig. A3 (chloride melt) and Fig. A4 (hydrothermal fluids). The contribution of cation in
420 the second coordination sphere of Au to the Fourier Transform (FT) of calculated EXAFS signal
421 increases with increasing cation atomic number and concentration, but has a significant input only
422 on the spectrum of chloride melt (Fig. A3). In case of chloride fluids (Fig. A4) the contribution of

423 the distant-coordination-sphere cation calculated from the AIMD atomic configurations is
424 underestimated relatively to the experimental spectra fitting. Nonetheless, the AIMD-EXAFS
425 calculations enabled us to explore the cation contribution to the EXAFS spectra using wavelet
426 transform (WT) diagrams.

427 Figure 6a shows contributions of scattering atoms (Cl, O, Cs) and multiple scattering
428 contribution Cl-Au-Cl to the WT of calculated AIMD-EXAFS spectra. The maxima of Cl
429 contribution to the 1st coordination sphere is located at $k \sim 6 \text{ \AA}^{-1}$, the Cl-Au-Cl multiple scattering
430 contribution is located at $k \sim 5.5 \text{ \AA}^{-1}$, and the maxima of O contribution is at $k \sim 4.5 \text{ \AA}^{-1}$ (the other
431 local maxima of the WT of EXAFS in Fig. 6a correspond to the distant coordination sphere atoms
432 and multiple scattering paths). The contribution of Cs is clearly seen in Fig. 6b which shows the
433 WT of EXAFS signal of experiment cap118 – the high-temperature fluid of relatively low density
434 ($575 \text{ }^\circ\text{C}/2100 \text{ bar}$, $d = 0.68 \text{ g}\cdot\text{cm}^{-3}$). Here the contribution of Cs in the second coordination sphere
435 occurs at $k \sim 9.5 \text{ \AA}^{-1}/R \sim 3 \text{ \AA}$. The WT of EXAFS of fluid of the same composition at higher
436 temperature and pressure (experiment cap 117, $450 \text{ }^\circ\text{C}/4500 \text{ bar}$, $d = 1.04 \text{ g}\cdot\text{cm}^{-3}$) does not exhibit a
437 notably heavy atom contribution except, probably, a small increase of the signal at $k \sim 9.7 \text{ \AA}^{-1}/R \sim$
438 4.5 \AA . The observed decrease of the heavy atom contribution to EXAFS signal, which takes place
439 with increase of the fluid density, is in line with experimental EXAFS spectra fittings performed
440 using ARTEMIS program (section 3.2, Table 4).

441

442 **3.5. AIMD-XANES spectra**

443 The Au L_3 -edge HERFD-XANES spectra calculated using atomic configurations obtained
444 by means of AIMD simulations are shown in Fig. 7. We calculated only the spectra of
445 CsCl/NaCl/KCl melt, HCl- and CsCl-dominated fluids in order to reveal, on a quantitative level, the
446 effect of the cation type (the lightest H^+ vs. the heaviest Cs^+) and cation concentration on the shape
447 of the HERFD-XANES spectra. It can be seen in the figure that the WL intensity depends on the
448 system composition. The most intense WLs correspond to systems with low salt concentration or

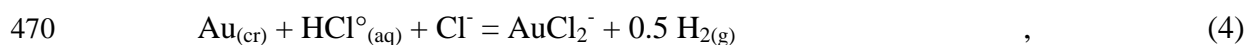
449 pure HCl fluid (spectra AuCl₂Cs and AuCl₂H, correspondingly). The WL intensity of 4m CsCl and
 450 CsCl/NaCl/KCl melt (spectra AuCl₂Cs₄ and melt, correspondingly) is lower. The decrease of the
 451 WL intensity in salt-rich systems is similar to the one observed in our HERFD-XAS experiments
 452 discussed in section 3.1. At the same time, the shift of the e.j. and the WL to lower energies in salt-
 453 rich systems predicted by the simulations were not observed in the experiments. We have to note
 454 that the energy positions of the spectral features in the *ab initio* calculations can be subjected to
 455 substantial uncertainty.

456 The observed peculiarities of the WL intensity can be explained by the effect of cation on
 457 the electronic density distribution within AuCl₂⁻ complex. Figure 8a demonstrates that the
 458 probability of finding a second-coordination-shell cation of Au has a maximum at a distance of ca.
 459 4 Å at Cs-Au-Cl angle of 65°. As an alkali metal cation approaches Au the positive charge of Au
 460 decreases (Fig. 8b) due to the decrease of 5*d* unoccupied DOS. This behavior of Au charge state
 461 yields the WL intensity decrease. Thus, the presence of an alkali metal cation in the distant
 462 coordination sphere of Au accounts for the observed decrease of the WL intensity of Au L₃-edge
 463 HERFD-XANES spectra of salt-rich systems.

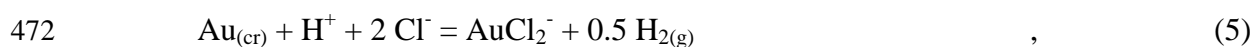
464

465 3.6. Speciation and concentration of Au in natural chloride fluids and melts

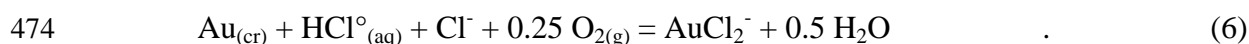
466 Salinity (chlorinity) of natural hydrothermal fluids varies in wide limits from almost dry
 467 melts to diluted aqueous solutions. Results of our study demonstrate that, independently of fluid
 468 chemistry (cation type and chloride salt concentration), AuCl₂⁻ is the dominant Au species in fluids
 469 of moderate to high density. The Au dissolution in aqueous fluids can be described by the reaction,



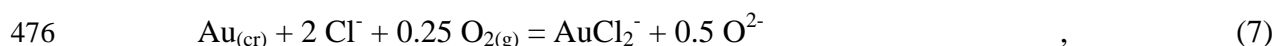
471 or, involving the proton activity (or pH = -log *a*(H⁺)),



473 or, using oxygen as a reactant which acts as an oxidizing agent,



475 Similar reaction can be written for Au dissolution in dry chloride melt,



477 or, in terms of neutral species,



479 These reactions imply that the solubility (concentration) of Au is governed by (i) the oxygen
480 fugacity, (ii) the fluid acidity, and, (iii) the chloride concentration. It follows from Eq. (4-8) that the
481 dissolved Au concentration in equilibrium with native gold increases proportionally to the first
482 power of fluid acidity, the second power of chloride salt concentration, and decreases as a square
483 root of hydrogen fugacity (or increases proportionally to $f(\text{O}_2)^{0.25}$). Predominance field of AuCl_2^-
484 falls within acidic range of pH. In sulfide-rich fluids the Au solubility is controlled by hydrosulfide
485 complex $\text{Au}(\text{HS})_2^-$ in the pH range from weakly acidic to alkaline, even in the presence of chlorides
486 (e.g., Renders and Seward, 1989; Shenberger and Barnes, 1989; Benning and Seward, 1996;
487 Stefánsson and Seward, 2003; Barnes, 2015; Trigub et al., 2017, see also references cited in these
488 papers), or by hydroxide complex $\text{Au}(\text{OH})^\circ$ at high (> 400 °C) temperatures (Zotov et al., 1985;
489 Stefánsson and Seward, 2003). The dissolved Au concentration in dry chloride melt is mostly
490 controlled by the oxygen fugacity and increases as a linear function of $f(\text{O}_2)^{0.25}$ (Eq. 7, 8).

491 Results of our spectroscopic experiments revealed the difference between the first and
492 second coordination spheres of Au complexes. The Cl^- ions in the nearest coordination sphere of Au
493 are bonded to the central atom via the formation of strong chemical bond whose nature can be
494 described in terms of interaction between the metal and ligand atomic orbitals with the formation of
495 molecular orbitals (cf. Figgs and Fitchman, 2000). In contrast, cations (H^+ , Na^+ , K^+ , Cs^+ in the case
496 of the present study) are weakly bonded to the negatively charged AuCl_2^- complex in the distant
497 coordination sphere of Au mostly via the electrostatic interaction. The difference in the chemical
498 bond properties is clearly demonstrated by the contrasting behavior of the atoms in the nearest and
499 distant coordination spheres depending on the experimental *TP*-parameters. Increasing pressure
500 results in weak contraction of the first coordination sphere of Au, whereas conjugated increase of

501 the dielectric constant of hydrothermal fluid causes expansion of the distant coordination sphere
502 composed of alkali metal cations. The number of cations in the distant coordination sphere of Au
503 also depends upon pressure, cation type, and salt concentration. From the point of view of
504 thermodynamics this means that the dissolution of Au can be explicitly described by the equilibrium
505 constants of chemical reactions which involve the formation of AuCl_2^- (Eq. 4-8), taking into
506 account the influence of weak outer sphere coulomb interactions via the activity coefficients. This
507 approach is valid for fluids with density of $d > 0.5 \text{ g}\cdot\text{cm}^{-3}$. In low-density fluids ($d < 0.3 \text{ g}\cdot\text{cm}^{-3}$),
508 where interaction between charged species becomes much more pronounced and neutral species
509 predominate, formation of ionic associates of the type of MeAuCl_2° and HAuCl_2° is hampered by
510 the change of predominant Au complex from charged AuCl_2^- to neutral AuCl° . The same
511 conclusion might be done with regard to other 1st group *d*-elements (Cu, Ag) which form stable
512 chloride complexes of similar composition. However, determination of composition and stability of
513 chloride complexes of these metals can only be done via accurate solubility measurements.

514 The Au solubilities monitored in the present study are in close agreement with
515 thermodynamic model of Akinfiev and Zotov (2010), which is based on the consideration of AuCl_2^-
516 complex. We note, however, that there is still a problem of disagreement between the values of
517 AuCl_2^- stability constants reported in the key studies on hydrothermal Au-Cl complexing of Zotov
518 et al. (1991) and Stefánsson and Seward (2003). The Au solubility constants obtained in Zotov et al.
519 (1991) are more than one log unit higher than those determined in Stefánsson and Seward (2003)
520 for similar *TP*-parameters. The reason of this deviation is unclear. This problem can be resolved by
521 additional measurements of Au solubility with rigorous in situ control of the most important
522 physicochemical parameters (e.g., $f(\text{H}_2)$), and in situ experimental solution sampling. This is an
523 ongoing study in our laboratory.

524

525

4. CONCLUSIONS

526 In the present study we recorded Au L_3 -edge X-ray absorption spectra of aqueous chloride-
527 bearing fluids at temperatures 350 – 575 °C and pressures up to 4500 bar. The composition of the
528 experimental fluids varied from HCl – dominated solutions to concentrated KCl and CsCl brines.
529 X-ray absorption spectra of dry CsCl/NaCl/KCl melt were recorded at 650 °C. Interpretation of
530 EXAFS spectra showed that, independently of the experimental fluid composition, AuCl_2^- complex
531 dominates the dissolved Au speciation. The distance between Au and the nearest-neighbor Cl atoms
532 varies between 2.25 – 2.28 Å. This interatomic distance is pressure-dependent and decreases with
533 increasing pressure from 2.28 Å (dry melt) to 2.25 Å (7.3m HCl/4.2m CsCl fluid at 4500 bar). The
534 outer coordination sphere of Au is composed of cations located at a distance of 3.3 – 4.1 Å. In
535 contrast to the inner coordination sphere, the outer coordination shell expands with increasing the
536 fluid pressure (dielectric constant). The number of cations located in the outer coordination sphere
537 increases with increasing cation atomic number and concentration. Such a contrasting behavior of
538 atoms located in the nearest and distant coordination spheres of Au reveals the difference in the
539 nature of the chemical bonds. Chloride ions are connected to Au by means of strong chemical
540 bonding, whereas weak coulomb interactions between AuCl_2^- and cations are responsible for the
541 formation of the disordered distant coordination sphere. The presence of cations in the distant
542 coordination sphere of Au is confirmed by the change of the shape of Au L_3 -edge HERFD-XANES
543 spectra in salt-rich brine and dry melt with respect to the HCl-dominated fluid. The results of our
544 study imply that the Au solubility in hydrothermal fluids in the wide range of TP -conditions and
545 chlorides concentration can be described in terms of the formation of AuCl_2^- complex using an
546 appropriate model for the activity coefficients.

547

548 **Acknowledgements**

549 The authors acknowledge the ESRF for the beamtime allocation under proposals 20-01-782
550 (BM20 beamline) and ES-360 (ID26 beamline). We thank Hugo Vitoux for technical support
551 during the in-situ experiment with a micro-tomo furnace at ID26, and Joerg Exner for technical

552 support at BM20. The authors are grateful to Alexander Zotov and Nikolai Akinfiev for useful
553 discussions during the manuscript preparation. Pavel Selivanov is acknowledged for preparation of
554 capillaries and technical assistance during the synchrotron experiments. We thank Max Wilke,
555 Elena Bazarkina, and Laurent Truche for providing us with silica glass capillaries. The results of
556 this work were obtained using the computational resources of MCC NRC “Kurchatov Institute”
557 (<http://computing.kiae.ru/>). Chemical analyses were carried out at the “IGEM-Analytica” Center for
558 collective use. This study was supported by the Russian Science Foundation grant No. 17-17-01220
559 (spectroscopic experiment and interpretation, AIMD calculations) and Basis theme RAS No. 72-7
560 (thermodynamic speciation calculations). The work of DACH is supported by the program 211 of
561 the Russian Federation Government, agreement No. 02.A03.21.0006 and by the Russian
562 Government Program of Competitive Growth of Kazan Federal University.

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Table 1. Compositions of experimental systems, experimental Au solubilities, and equilibrium concentrations of aqueous species calculated using thermodynamic data available in the literature: AuCl₂⁻ - Akinfiyev and Zotov (2010), CsCl^o - Sverjensky et al. (1997), HCl^o - Tagirov et al. (1997), other aqueous species, H₂O, and Au_(cr) – SUPCRT92 (Johnson et al., 1992).

Capillary No.	28	30	110	110	117	118	122	122	122	122
Solute Concentr.	3.8mHCl/ 0.02 mNaCl/ 0.5mH ₂ SO ₄ / 0.5m H ₂ SO ₃	3.4mNaCl/ 0.7mH ₂ SO ₄ / 0.6m H ₂ SO ₃	3.8mHCl/ 0.57mKCl/ 0.03mK ₂ S ₂ O ₈	3.8mHCl/ 0.57mKCl/ 0.03mK ₂ S ₂ O ₈	7.3mHCl/ 4.2mCsCl/ 0.56mH ₂ SO ₄ / 0.12mNa ₂ SO ₃	7.3mHCl/ 4.2mCsCl/ 0.56mH ₂ SO ₄ / 0.12mNa ₂ SO ₃	3.8mHCl/ 7.2mKCl/ 0.4mK ₂ S ₂ O ₈	3.8mHCl/ 7.2mKCl/ 0.4mK ₂ S ₂ O ₈	3.8mHCl/ 7.2mKCl/ 0.4mK ₂ S ₂ O ₈	3.8mHCl/ 7.2mKCl/ 0.4mK ₂ S ₂ O ₈
<i>d</i> , g·cm ⁻³	0.53	0.67	0.71	0.71	1.04	0.68	0.96	0.96	0.96	0.96
<i>t</i> , °C/ <i>P</i> , bar	485 °C/ 600 bar	485 °C/ 700 bar	400 °C/ 850 bar	450 °C/ 2100 bar	450 °C/ 4500 bar	575 °C/ 2100 bar	350 °C/ 150 bar	400 °C/ 500 bar	450 °C/ 1100 bar	525 °C/ 1900 bar
Calculated concentrations, mol·(kg H ₂ O) ⁻¹										
H ⁺	5.73E-01	1.43E-02	4.43E-01	4.55E-01	2.21E+00	6.83E-02	3.01E-01	1.89E-01	1.04E-01	3.52E-02
OH ⁻	1.03E-10	7.10E-09	2.72E-10	8.84E-10	1.44E-09	5.92E-09	1.93E-10	9.86E-10	2.87E-09	1.53E-08
O ₂	5.33E-10	1.62E-14	9.74E-19	1.05E-17	1.11E-19	4.37E-13	1.43E-16	6.73E-16	2.89E-15	1.48E-13
H ₂	3.48E-09	2.62E-07	8.69E-09	8.01E-08	2.91E-07	3.07E-06	1.53E-11	6.29E-10	1.10E-08	2.25E-07
HSO ₃ ⁻	8.89E-09	1.73E-06	1.90E-08	7.52E-08	1.40E-06	9.07E-07	8.91E-10	2.07E-08	2.34E-07	1.83E-06
HSO ₄ ⁻	3.94E-01	6.69E-01	4.04E-02	2.55E-02	9.33E-02	3.81E-02	7.43E-01	7.70E-01	7.47E-01	6.00E-01
SO ₂ ^o	5.82E-01	6.15E-01	5.62E-03	2.04E-02	2.52E-01	6.27E-01	1.09E-04	2.91E-03	2.77E-02	1.86E-01
SO ₄ ²⁻	8.30E-06	5.27E-04	8.20E-07	3.33E-07	1.24E-06	7.77E-07	1.18E-04	2.06E-04	1.13E-04	8.59E-05
H ₂ SO ₄ ^o	1.20E-02	8.74E-04	1.39E-02	1.40E-02	3.30E-01	1.17E-03	5.63E-02	2.70E-02	2.42E-02	8.12E-03
Cl ⁻	1.79E-02	3.15E-01	6.96E-01	7.55E-01	3.07E+00	6.03E-01	1.61E+00	1.93E+00	1.91E+00	1.75E+00
HCl ^o	3.40E+00	6.11E-01	3.27E+00	3.21E+00	4.63E+00	5.87E+00	2.64E+00	2.77E+00	2.79E+00	2.38E+00
Na ⁺	1.49E-02	1.01E+00			1.27E-01	9.45E-02				
NaCl ^o	4.89E-03	2.35E+00			1.11E-01	1.41E-01				
K ⁺			3.65E-01	4.27E-01			2.85E+00	3.32E+00	3.40E+00	3.49E+00
KCl ^o			2.65E-01	2.03E-01			5.15E+00	4.68E+00	4.59E+00	4.46E+00
Cs ⁺					1.09E+00	1.50E+00				
CsCl ^o					3.08E+00	2.62E+00				
AuCl ₂ ⁻	1.76E-01	4.39E-02	7.12E-02	1.01E-01	2.65E-01	1.02E+00	8.00E-01	8.06E-01	8.55E-01	1.17
Au _{exp.} ^(a)	n/d	n/d	3E-02	5E-02	2E-01	7E-01	2E-01	2E-01	2E-01	2E-01
<i>I</i> ^(b)	0.588	1.029	0.808	0.881	3.426	1.66	3.153	3.507	3.507	3.522
pH	1.386	2.864	0.881	0.797	0.154	1.87	1.366	1.586	1.753	2.19
Eh	0.34	-0.039	0.28	0.235	0.255	0.017	0.378	0.279	0.186	0.042
^a dissolved Au concentration determined from the amplitude of the absorption edge height of Au L ₃ -edge transmission spectra, estimated uncertainty ±0.3 log units; ^b ionic strength.										

Table 2. Results of *ab initio* molecular dynamics (AIMD) simulations performed for systems containing Au complexes of different compositions. Uncertainties correspond to the standard deviations.

AIMD calculation	Composition of simulation box	Box size, Å	Starting Au species ^a	T, K	Density ρ , g/cm ³ (without Au)	Density ρ , g/cm ³ (with Au)	Simulation time, ps
A	Au ⁺ , H ⁺ , 2Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	722 (46)	0.55	0.65	23
B	Au ⁺ , Na ⁺ , 2Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	722 (45)	0.56	0.67	23
C	Au ⁺ , K ⁺ , 2Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	722 (45)	0.57	0.67	25
D	Au ⁺ , Cs ⁺ , 2Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	722 (46)	0.62	0.72	23
E	Au ⁺ , 4Cs ⁺ , 5Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	722 (48)	0.83	0.98	24
F	Au ⁺ , 48Cs ⁺ , 32Na ⁺ , 27K ⁺ , 5Cl ⁻	16.3		923 (53)	2.13	2.17	15
G	Au ⁺ , 4Cs ⁺ , 5Cl ⁻ , 55H ₂ O	14.7	Au(Cl) ₂ ⁻	723 (49)	0.83	0.98	15

^a starting system configurations were prepared using GROMACS software and included the Au complexes of indicated composition (see section 2.4 for calculation details). Starting geometry for F calculation was prepared using CsCl crystal lattice by substituting Cs atoms by Au, Na and K atoms, and correction of lattice constant to get the specified density.

Table 3. Position of edge jump (e.j.), white line (WL), and the next to WL feature (B) of Au L_3 -edge HERFD- and TFY-XANES spectra (as determined by IFEFFIT program). Uncertainty of the energy values is ± 0.5 eV.

Experimental system	Feature	Position, eV
HERFD-XANES (± 0.2 eV)		
cap 5407 Melt CsCl/NaCl/KCl $t = 650$ °C 45.5at.%/30 at%/24.5 at%	e.j.	11919.5
	WL	11921.0
	B	11930.8
	ΔB -WL ^a	9.8
cap 28 $t = 485$ °C, $P = 600$ bar $d = 0.53$ g·cm ⁻³ 3.8m HCl/0.02 m NaCl 0.5m H ₂ SO ₄ /0.5m H ₂ SO ₃	e.j.	11919.3
	WL	11920.7
	B	11931
	ΔB -WL ^a	10.3
cap 30 $t = 485$ °C, $P = 700$ bar $d = 0.67$ g·cm ⁻³ 3.4 m NaCl 0.7m H ₂ SO ₄ /0.6m H ₂ SO ₃	e.j.	11919.6
	WL	11920.8
	B	11930.5
	ΔB -WL ^a	9.8
TFY-XANES (± 0.5 eV)		
cap 2156 Melt CsCl/NaCl/KCl 45.5at.%/30 at%/24.5 at% $t = 650$ °C	e.j.	11919.5
	WL	11922.5
	B	11935.5
cap 2156 Crystallized melt CsCl/NaCl/KCl Ambient temperature	e.j.	11918
	WL	11920
	B	11934
Cap110 $t = 400$ °C, $P = 850$ bar $d = 0.71$ g·cm ⁻³ 3.8m HCl/0.57m KCl/ 0.03 m K ₂ S ₂ O ₈	e.j.	11919
	WL	11922
	B	11931/11933
Cap110 $t = 450$ °C, $P = 2100$ bar	e.j.	11919
	WL	11922
	B	11931/11933
Cap117 $t = 450$ °C, $P = 4500$ bar $d = 1.04$ g·cm ⁻³ 7.3m HCl/4.2m CsCl 0.56m H ₂ SO ₄ /0.12m Na ₂ SO ₃	e.j.	11919
	WL	11922
	B	11931/11933
Cap118 $t = 575$ °C, $P = 2100$ bar $d = 0.68$ g·cm ⁻³ 7.3m HCl/4.2m CsCl 0.56m H ₂ SO ₄ /0.12m Na ₂ SO ₃	e.j.	11919
	WL	11922
	B	11932
Cap122 $t = 350$ °C, $P = 150$ bar $d = 0.96$ g·cm ⁻³ , 3.8m HCl/7.2m KCl/ 0.4m K ₂ S ₂ O ₈	e.j.	11918.5
	WL	11921.5
	B	11931/11933
Cap122 $t = 400$ °C, $P = 500$ bar	e.j.	11918.5
	WL	11921.5
	B	11930/11933
Cap122 $t = 450$ °C, $P = 1100$ bar	e.j.	11918.5
	WL	11921.5
	B	11930/11933
Cap122 $t = 525$ °C, $P = 1900$ bar	e.j.	11918.5
	WL	11921.5
	B	11931

^a energy difference between maxima of B and WL features.

Table 4. Results of EXAFS spectra fitting (ARTEMIS program, $S_0^2 = 0.84$, AuCs₂Cl₆ was accepted as initial model structure, Matsushita and Ahsbahs, 1998). Experimental spectra fits were performed in k -space if not indicated otherwise.

Atom	EXAFS				
	N	$R, \text{\AA}$	$\sigma^2, \text{\AA}^{-2}$	Fit quality	
				E°, eV	Red. χ^2 (R-factor)
Cap 2156, melt CsCl/NaCl/KCl (k-space: 3 – 13), $t = 650 \text{ }^\circ\text{C}$					
Cl	2	2.28±0.007	0.004±0.001	8.8±1.0	(0.033)
Cs	0.19±0.21	4.03±0.05	0.002±0.001		
MS ¹	2	4.60±0.05	0.021±0.009		
Cap 2156, crystallized melt CsCl/NaCl/KCl (k-space: 3 – 13), ambient temperature					
Cl	1.88±0.32	2.29±0.009	0.001±0.001	8.0±1.7	(0.060)
Cl	0.96±0.79	2.94±0.03	0.003±0.006		
Cs	1	3.26±0.08	0.012±0.011		
MS ¹	1.88	4.58±0.05	0.004±0.007		
MS ¹	0.96±0.79	5.08±0.05	0.003±0.022		
Cap110 (fit in R-space, R-range: 1.3-4.4; k-range: 3 – 11), $t = 400 \text{ }^\circ\text{C}$, $P = 850 \text{ bar}$ $d = 0.71 \text{ g}\cdot\text{cm}^{-3}$, 3.8m HCl/0.57m KCl/0.03 m K₂S₂O₈					
Cl	2.27±0.13	2.27±0.005	0.006±0.001	7.5±2.5	0.5 (0.007)
K	1	4.15±0.02	0.003±0.002		
MS ¹	2.27±0.13	4.53±0.02	0.016±0.004		
Cap110 (fit in R-space, R-range: 1.3-4.4; k-range: 3 – 10), $t = 450 \text{ }^\circ\text{C}$, $P = 2100 \text{ bar}$					
Cl	2.08±0.23	2.26±0.009	0.004±0.001	7.2±1.2	3.0 (0.013)
K	1	4.12±0.07	0.004±0.06		
MS ¹	2.08±0.23	4.50±0.04	0.006±0.005		
Cap117 (fit in R-space, R-range: 1.3-4.4; k-range: 3 – 13), $t = 450 \text{ }^\circ\text{C}$, $P = 4500 \text{ bar}$ $d = 1.04 \text{ g}\cdot\text{cm}^{-3}$, 7.3m HCl/4.2m CsCl (+0.56m H₂SO₄/0.12m Na₂SO₃)					
Cl	2.02±0.13	2.25±0.005	0.004±0.001	4.6±0.8	1.3 (0.013)
Cs	0.59±0.86	4.11±0.03	0.005±0.007		
MS ¹	2.02±0.13	4.48±0.03	0.020±0.008		
Cap118 (fit in R-space, R-range: 1.3-4.4; k-range: 3 – 11), $t = 575 \text{ }^\circ\text{C}$, $P = 2100 \text{ bar}$ $d = 0.68 \text{ g}\cdot\text{cm}^{-3}$, 7.3m HCl/4.2m CsCl (+0.56m H₂SO₄/0.12m Na₂SO₃)					
Cl	1.76±0.26	2.27±0.01	0.003±0.001	6.6±1.6	13.7 (0.030)
Cs	0.30±0.63	3.31±0.05	0.003±0.013		
MS ¹	1.76±0.26	4.52±0.08	0.019±0.012		

Table 4 - continued

Cap122 (fit in R -space, R -range: 1.3-4.4; k -range: 3 – 13), $t = 350$ °C, $P = 150$ bar (possible presence of vapor bubble) $d = 0.96$ g·cm ⁻³ , 3.8m HCl/7.2m KCl/0.4m K ₂ S ₂ O ₈					
Cl	2.04±0.18	2.28±0.006	0.003±0.001	8.8±1.1	7.1 (0.022)
K	0.53±1.49	4.08±0.06	0.003±0.016		
MS ¹	2.04±0.18	4.54±0.08	0.018±0.009		
Cap122 (fit in R -space, R -range: 1.3-4.4; k -range: 3 – 10). $t = 400$ °C, $P = 500$ bar					
Cl	1.85±0.24	2.27±0.01	0.003±0.001	8.0±1.4	1.4 (0.026)
MS ¹	1.85±0.24	4.54±0.04	0.010±0.006		
Cap122 (k -range: 3 – 13). $t = 450$ °C, $P = 1100$ bar					
Cl	2.11±0.19	2.27±0.007	0.004±0.001	8.1±1.1	3.6 (0.024)
K	1	4.06±0.04	0.005±0.004		
MS ¹	2.11±0.19	4.49±0.04	0.016±0.008		
Cap122 (k -range: 3 – 12). $t = 525$ °C, $P = 1900$ bar					
Cl	1.93±0.19	2.27±0.008	0.004±0.001	7.5±1.2	3.9 (0.026)
K	1	3.97±0.02	0.014±0.016		
MS ¹	1.93±0.19	4.53±0.07	0.023±0.013		
¹ MS – multiple scattering, Au-Cl1-Au-Cl2.					

Table 5. Geometries of Au aqueous complexes obtained by the *ab initio* molecular dynamics (AIMD) simulations (simulation details are given in Table 2). Uncertainties correspond to the standard deviations. Cation coordination number N corresponds to number of cations inside a sphere with radius of 6 Å.

AIMD calculation	Fluid composition	Complex	Au-Cl, Å	Cl-Au-Cl, °	Cation N
A	2m HCl	Au(Cl) ₂ ⁻	2.35 (0.078)	167 (6.2)	-
B	2m NaCl	Au(Cl) ₂ ⁻	2.35 (0.065)	169 (5.6)	0.17
C	2m KCl	Au(Cl) ₂ ⁻	2.35 (0.074)	168 (5.8)	0.36
D	2m CsCl	Au(Cl) ₂ ⁻	2.35 (0.087)	169 (6.1)	0.41
E	4m CsCl	Au(Cl) ₂ ⁻	2.35 (0.078)	165 (8.0)	1.31
F	CsCl/NaCl/ KCl	Au(Cl) ₂ ⁻	2.40 (0.14)	159 (16.0)	8.32
G	4m CsCl	Au(Cl) ₂ ⁻	2.34 (0.076)	168 (7.0)	1.63

Figure captions

Fig. 1. Gold L_3 -edge HERFD-XANES spectra of Au-bearing solutions (Cap28, 30) and melt (Cap5407). Vertical dashed lines indicate the main spectral features.

Fig. 2. Gold L_3 -edge TFY-XANES spectra of: (a) - Au-bearing solutions (Cap110, 117, 118, 122) and melt (Cap2156), and (b) – melt (Cap2156) at 650 °C and after crystallization at ambient temperature.

Fig. 3. Comparison of experimental Au L_3 -edge EXAFS spectra (*left*) and their Fourier Transforms (*right*).

Fig. 4. Results of Au L_3 -edge EXAFS spectra fits performed by means of ARTEMIS program. (a) CsCl/KCl/NaCl melt (Cap2156). (b-d) Aqueous fluids cap110, cap117, 118, cap122. *Top* - normalized k^2 -weighted EXAFS spectra; *Bottom* - corresponding Fourier Transforms (not corrected for phase shift). Thick black line – experimental spectra, thin dotted red line – fit. Experimental TP -parameters and fluid compositions are indicated at the top of each figure.

Fig. 5. Atomic densities and coordination numbers of Au complexes obtained using *ab initio* molecular dynamics (AIMD) simulations (see Table 2 for system compositions).

Fig. 6. EXAFS signal and wavelet transform (WT) of the EXAFS signal (distance R (Å) vs. wavevector k (Å⁻¹)). (a) contributions of different scattering atoms to the EXAFS spectra, AIMD calculation E (4m CsCl). (b) experimental spectra cap117 and cap118. Not corrected for phase shift. The WT intensity scales are shown at the right side of each WT image.

Fig. 7. Gold L_3 -edge HERFD-XANES spectra calculated using atomic configurations generated via AIMD simulations. Spectrum AuCl₂H corresponds to calculation A, AuCl₂Cs – calculation D, AuCl₂Cs₄ – calculation E, melt – calculation F. Compositions of simulation boxes are given in Table 2.

Fig. 8. Position of Cs calculated by AIMD (calculation E) (a), and positive Bader partial atomic charge of Au as a function of Au-Cs distance (b). Each point in Fig. (b) corresponds to average AIMD configuration for ± 0.12 Å interval around the indicated Au-Cs distance.

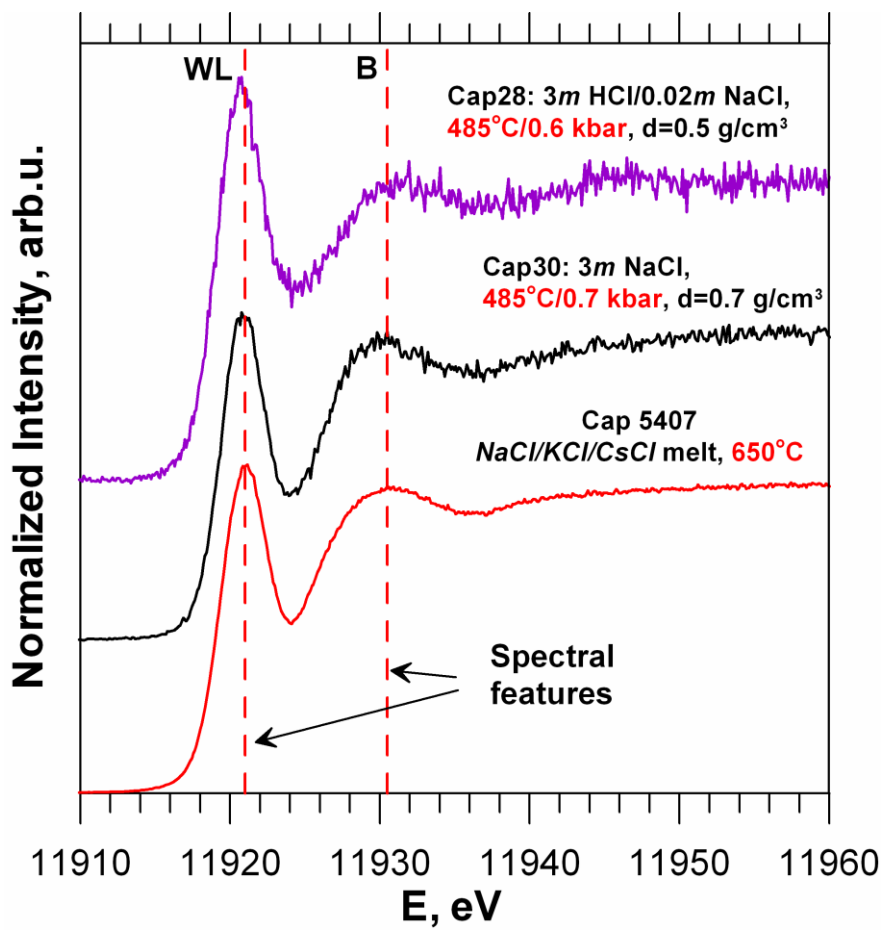


Fig. 1

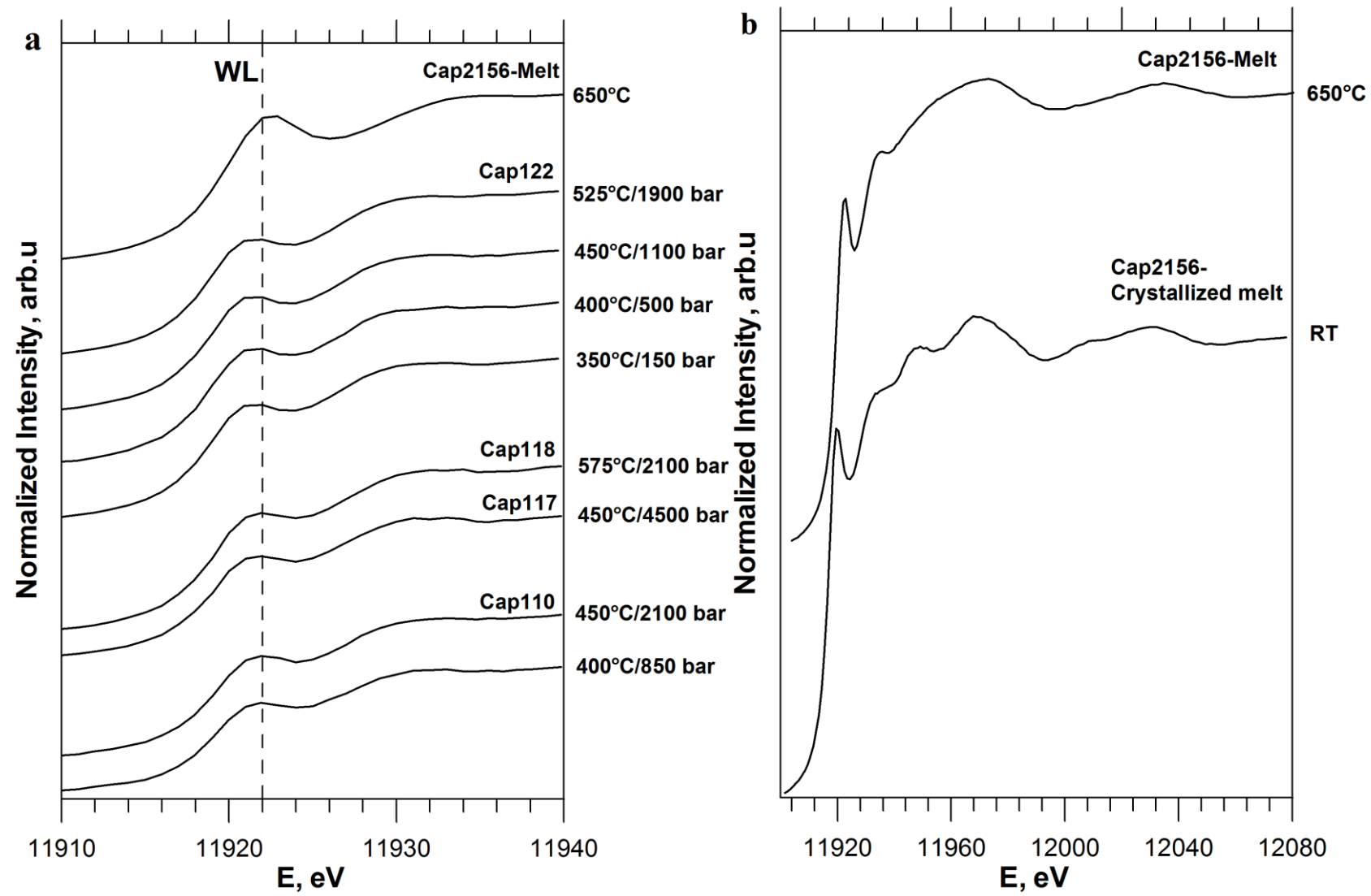


Fig. 2

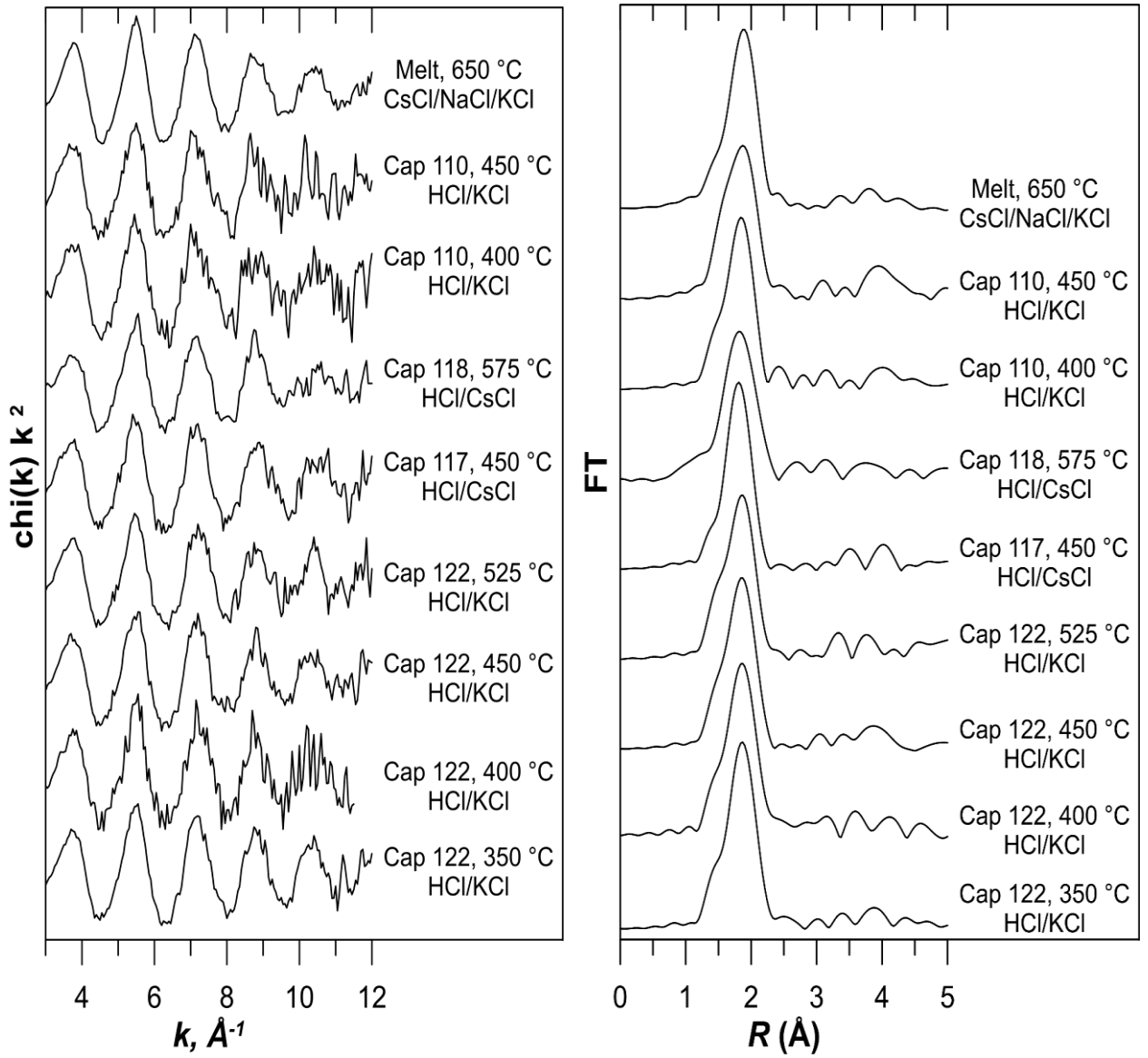


Fig. 3

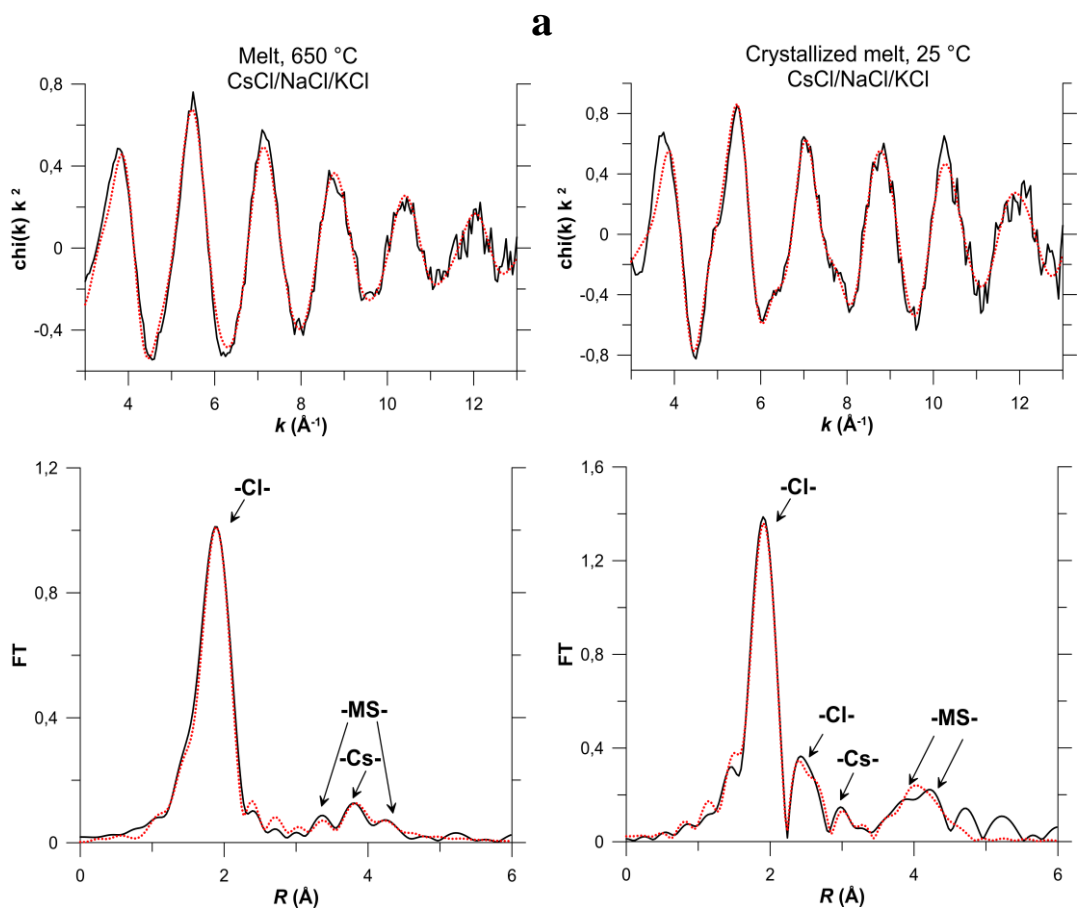


Fig. 4

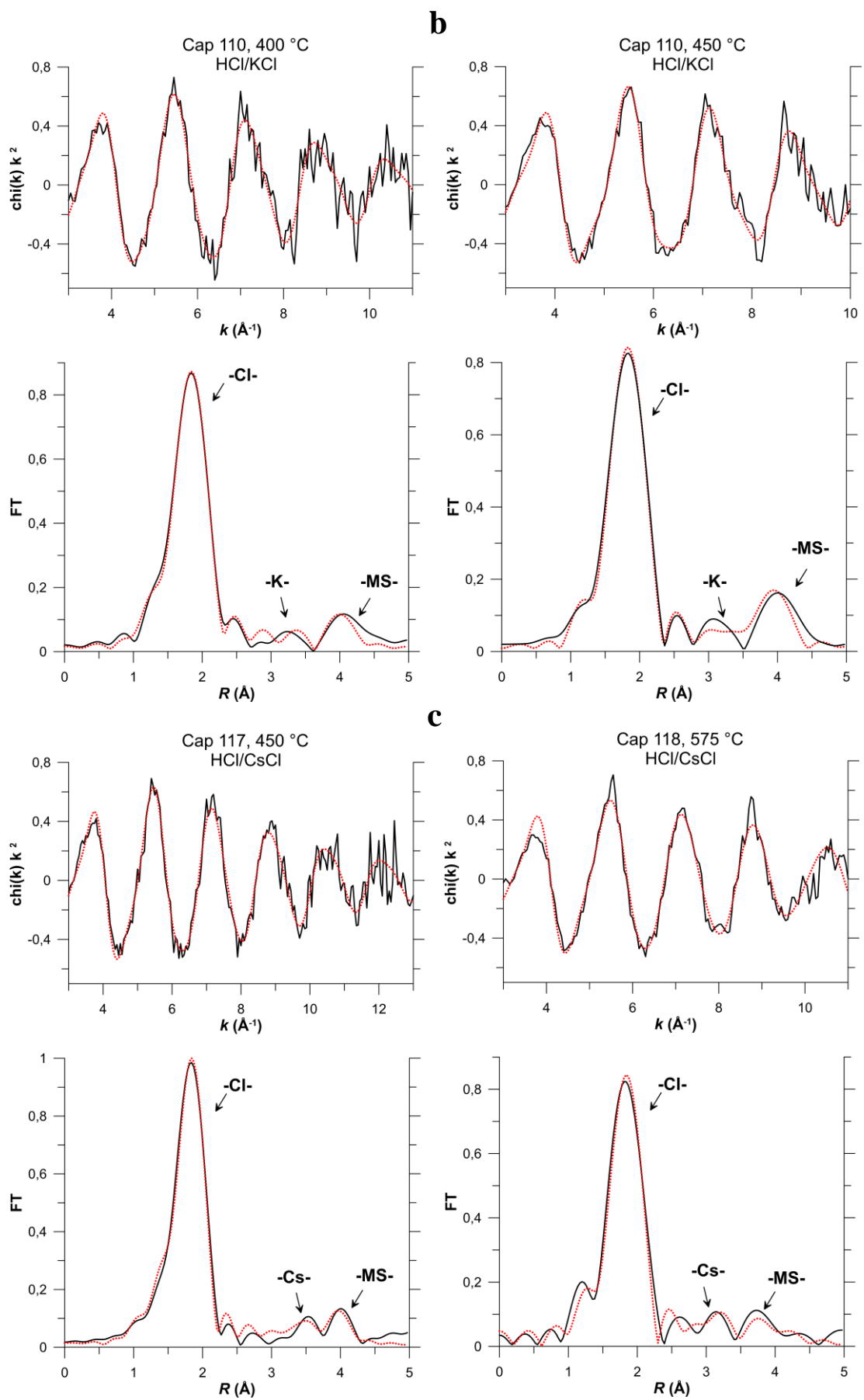


Fig. 4 - continued

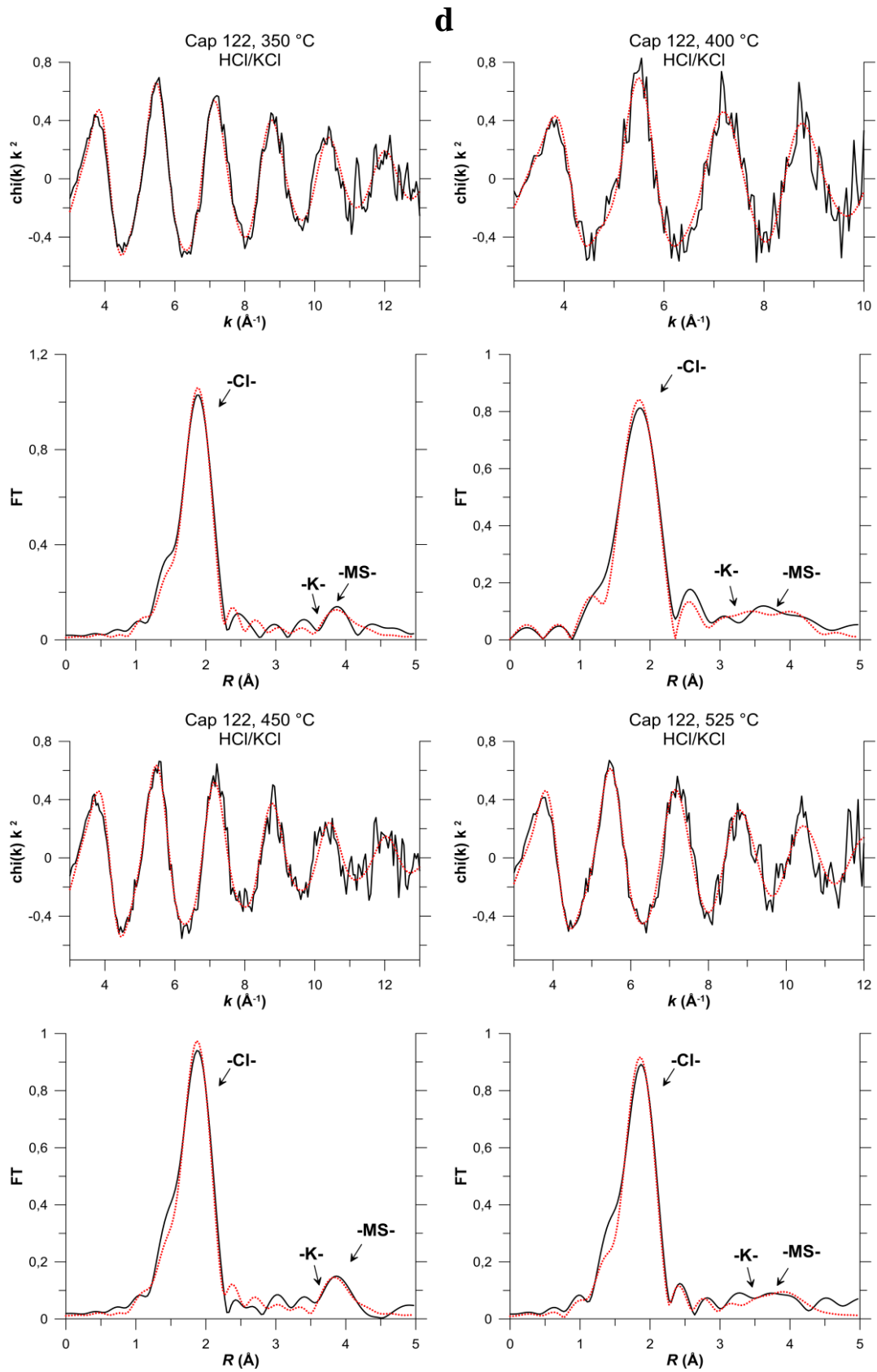


Fig. 4 – continued

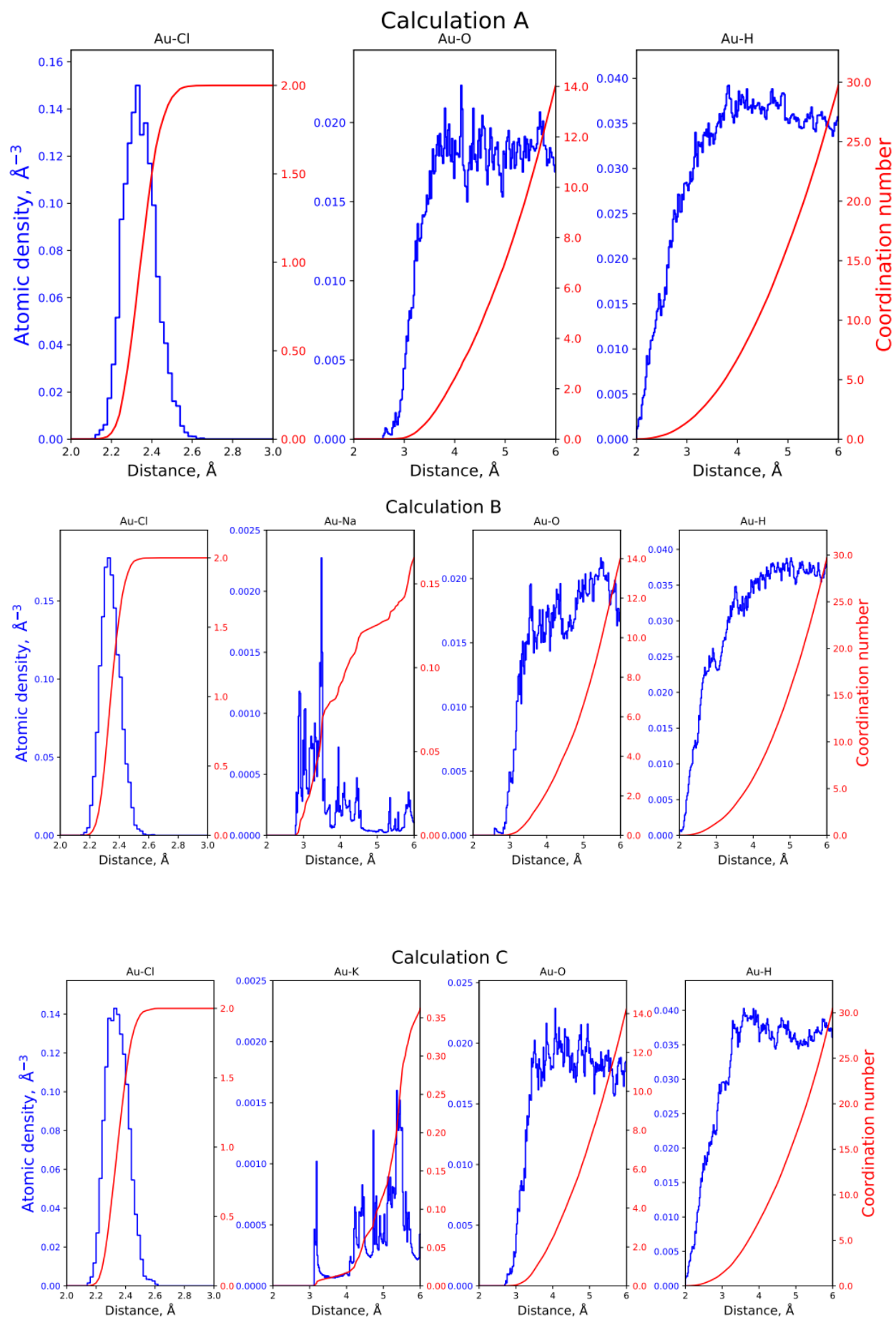


Fig. 5

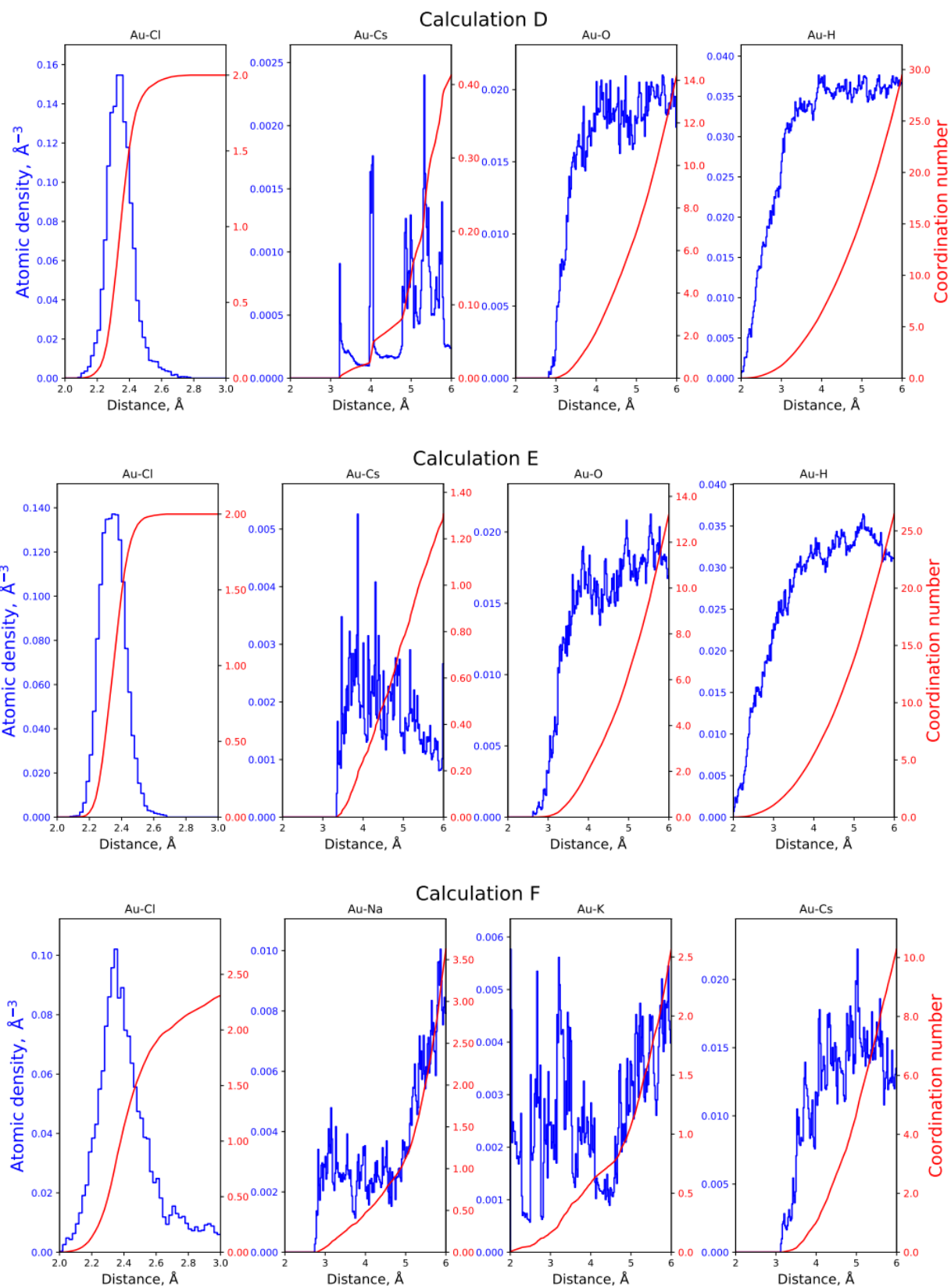


Fig. 5 – continued

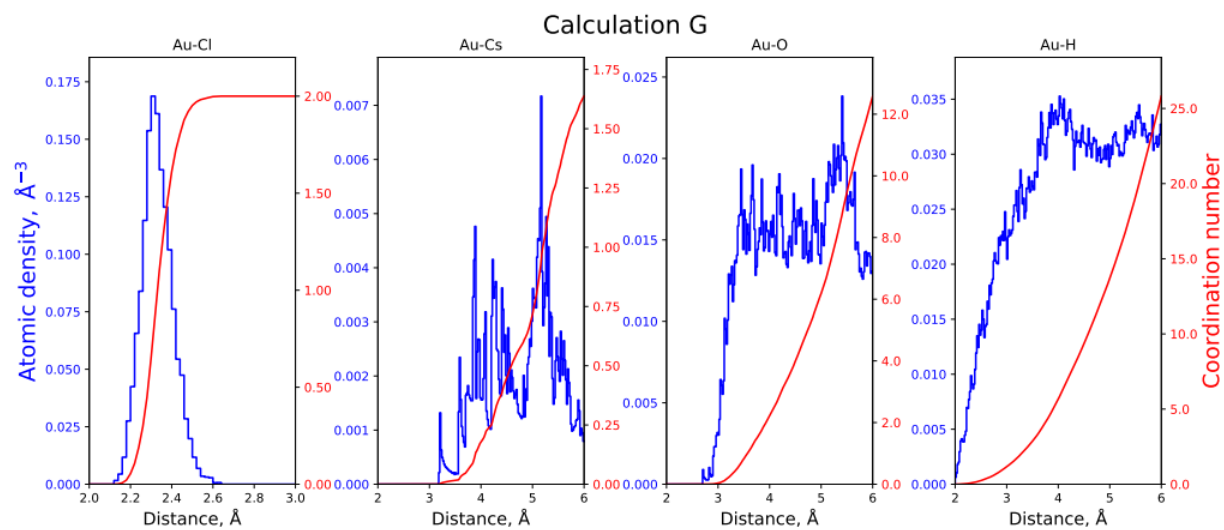


Fig. 5 – continued

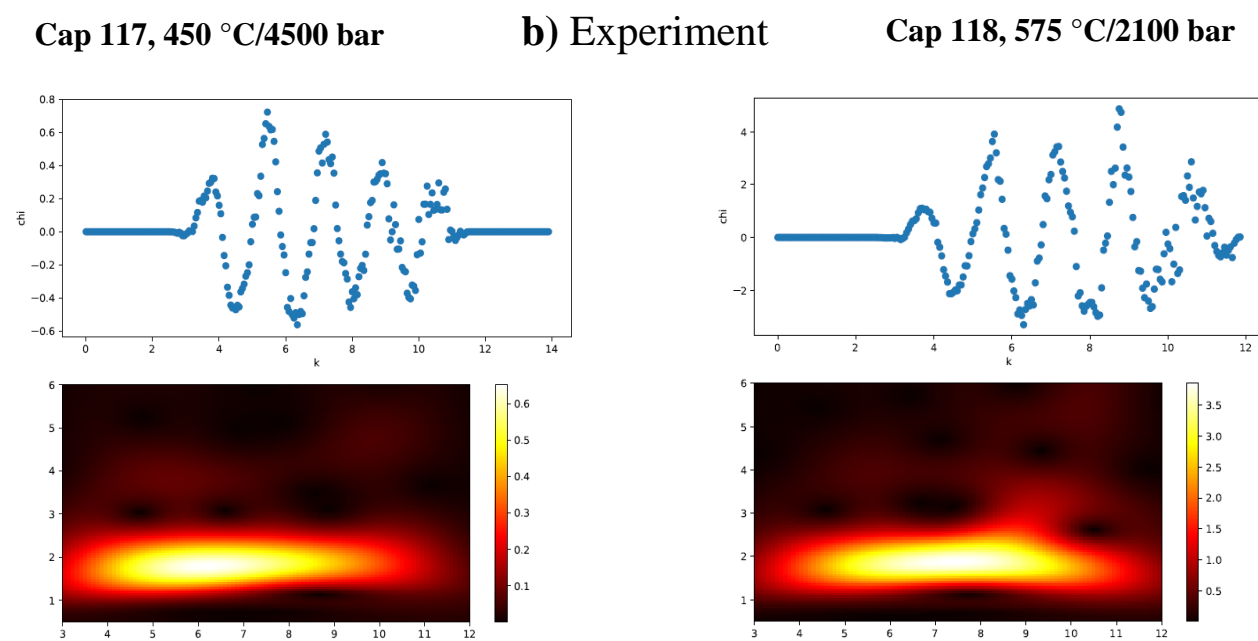
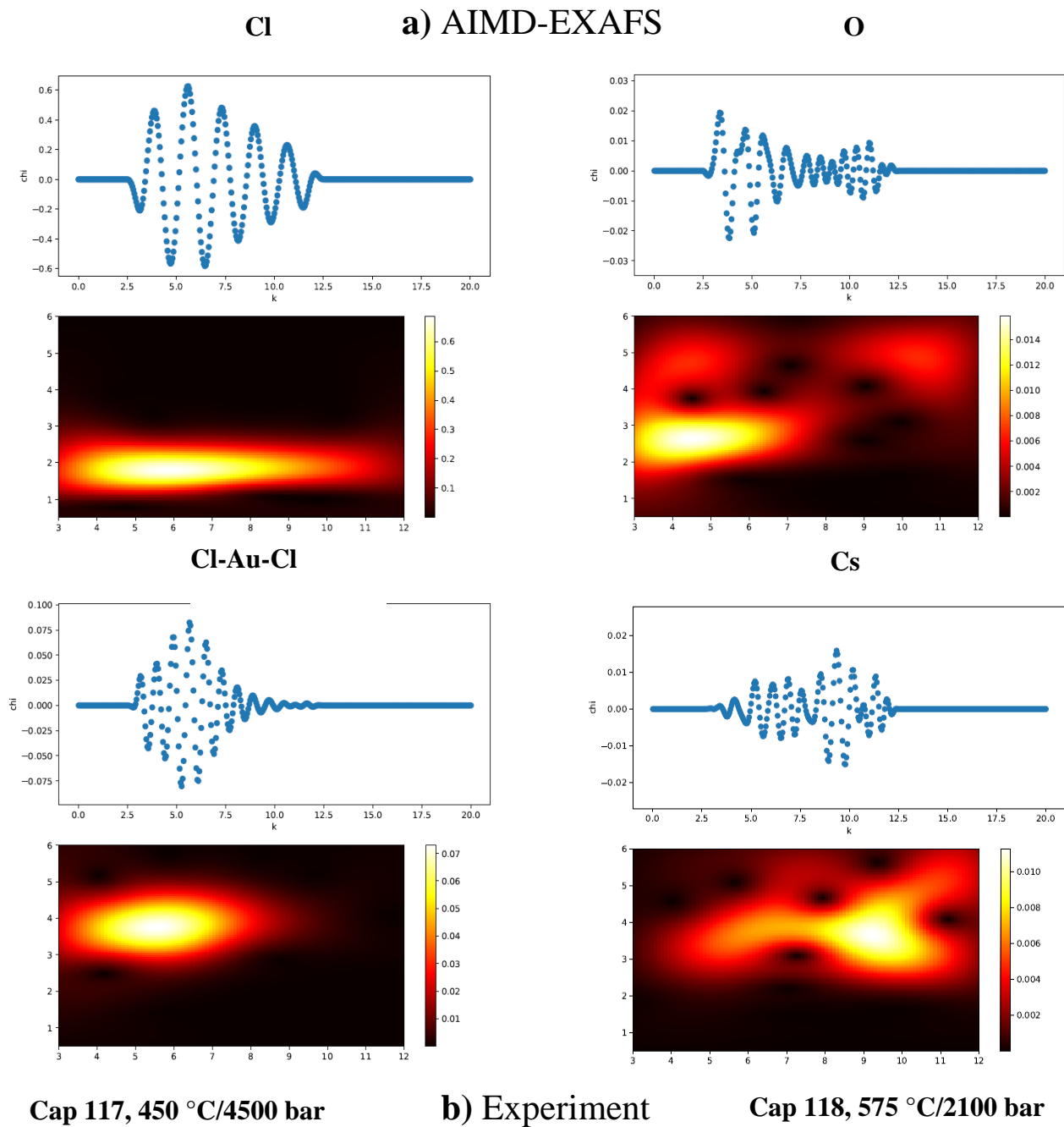


Fig. 6

Au L₃ XANES spectra

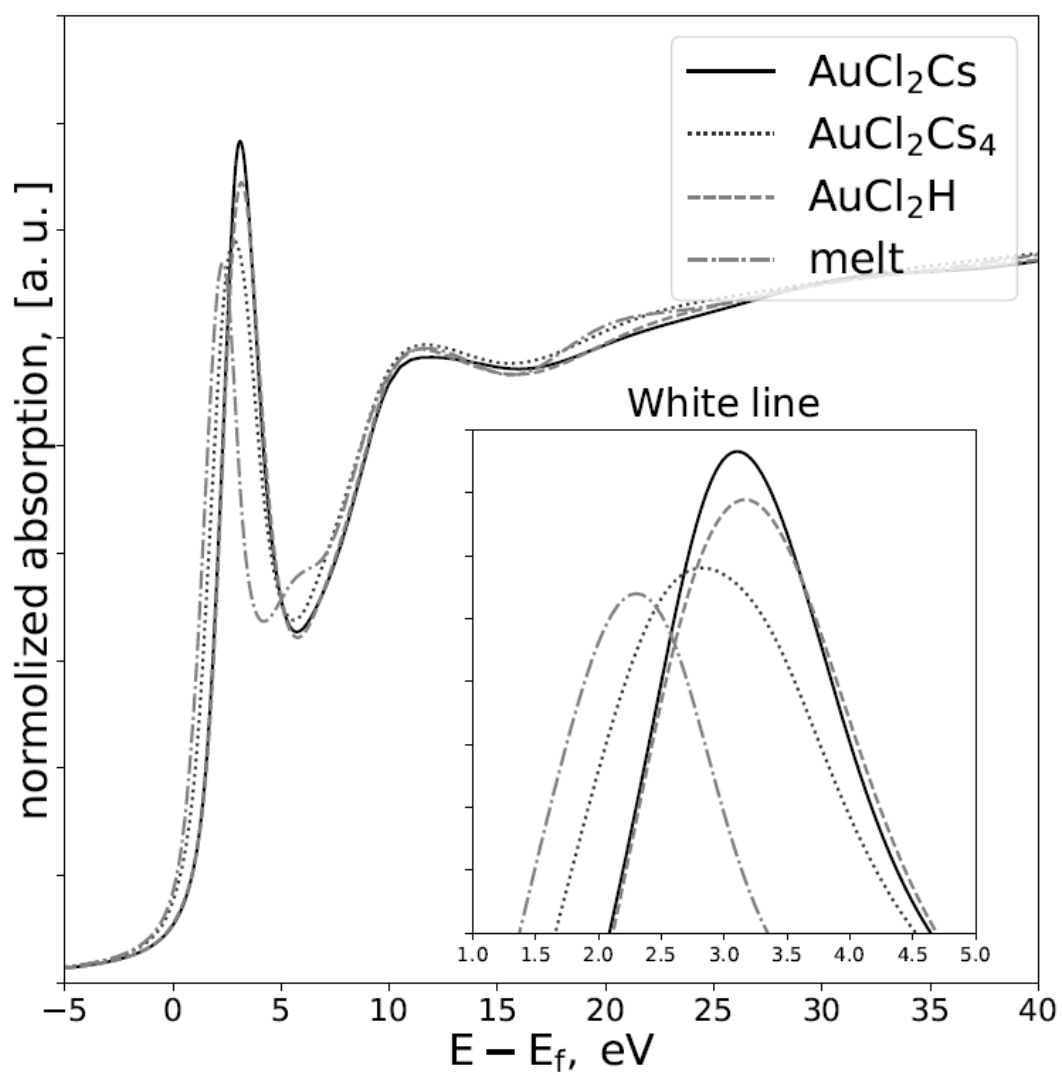


Fig. 7

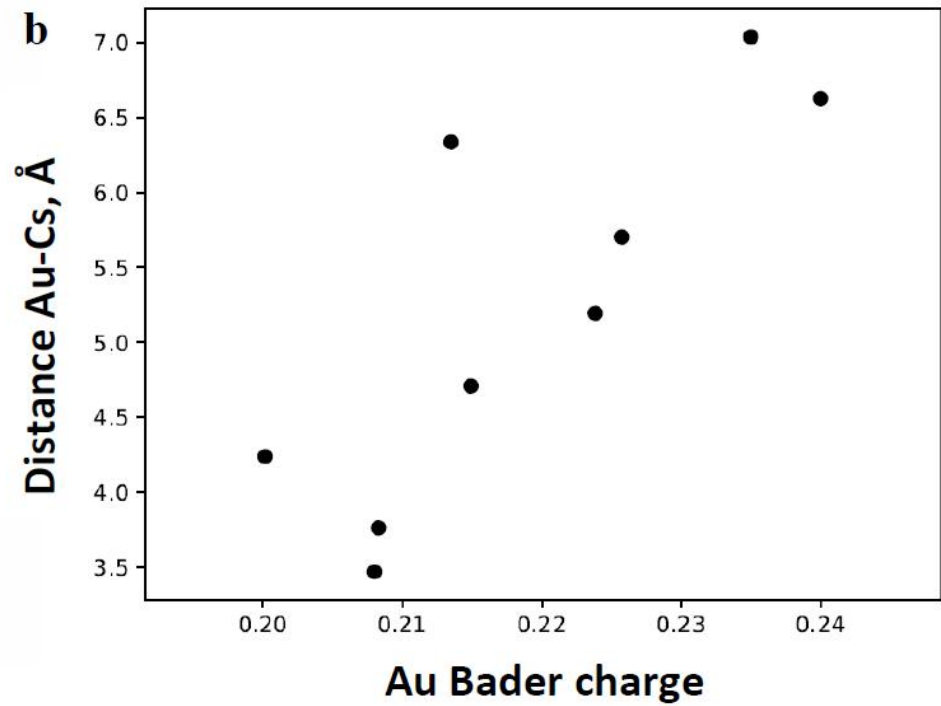
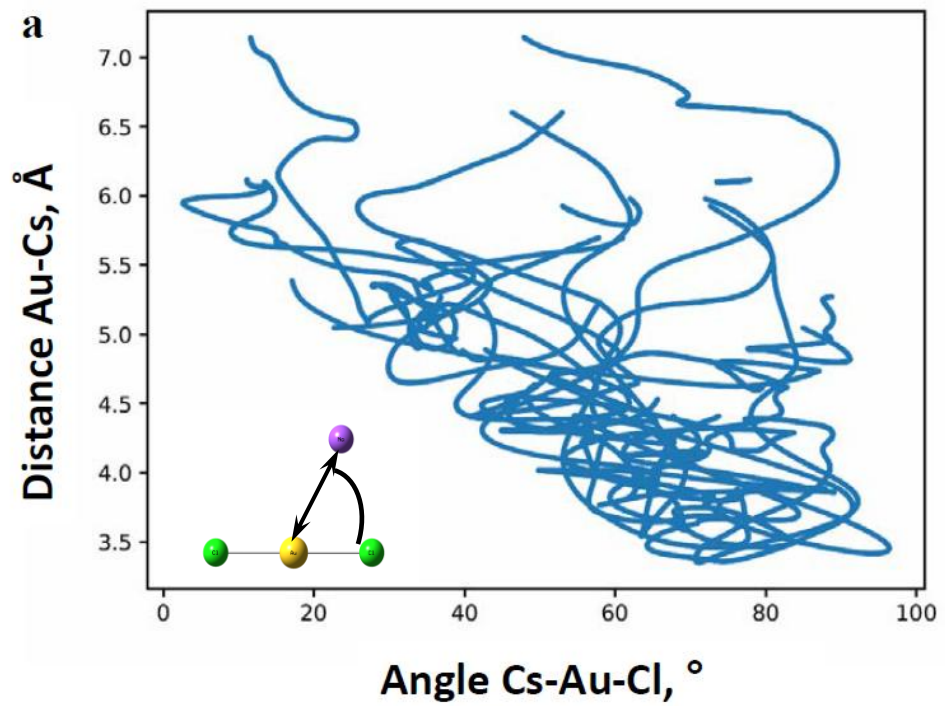


Fig. 8

Electronic Annex

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